

# Adsorption of Tetramethylammonium Hydroxide on Activated Carbon

Devarly Prahas<sup>1</sup>; J. C. Liu<sup>2</sup>; Suryadi Ismadji<sup>3</sup>; and Meng-Jiy Wang<sup>4</sup>

**Abstract:** Tetramethylammonium hydroxide (TMAH) is widely used in manufacturing thin-film transistor liquid crystal displays (TFT-LCD) and semiconductors. Because of its toxicity and nitrogen content, TMAH-containing wastewater has to be properly treated before discharge. In this study, several commercial activated carbons were characterized and used for TMAH adsorption. Activated carbon with micropores showed higher adsorption capacity for TMAH. The adsorption kinetics could be represented by pseudo-second-order model, and the Langmuir model fitted the adsorption of TMAH on activated carbon under different pH adsorption well. Higher pH was favorable for TMAH adsorption, while there was no TMAH adsorption at  $\text{pH } 4.7 \pm 0.2$ . The adsorption was spontaneous and exothermic. Exhausted carbon could be easily regenerated by 0.1 N HCl, with no considerable reduction in performance even after five cycles of adsorption-desorption. It was proposed that electrostatic interaction was the main mechanism of TMAH adsorption on the activated carbon. **DOI:** [10.1061/\(ASCE\)EE.1943-7870.0000482](https://doi.org/10.1061/(ASCE)EE.1943-7870.0000482). © 2012 American Society of Civil Engineers.

**CE Database subject headings:** Adsorption; Activated carbon; pH; Wastewater management.

**Author keywords:** Adsorption; Activated carbon; pH; Tetramethylammonium hydroxide (TMAH); Thin-film transistor liquid crystal display (TFT-LCD); Wastewater.

## Introduction

In Taiwan, thin-film transistor liquid crystal displays (TFT-LCD) and semiconductors are currently the most important industries. Tetramethylammonium hydroxide (TMAH) is a major ingredient of photoresist developer in the lithography step of the LCD manufacturing process (Sugawara et al. 2002). TMAH is used as an alkaline etchant and as a cleaning solution in semiconductor manufacturing (Thong et al. 1997). Consequently, a high amount of TMAH-containing wastewater from these industries can be expected. The typical TMAH concentration ranges from a few to hundreds of milligram per liter, depending upon the manufacturing process and its operation. TMAH is highly toxic and several symptoms of poisoning caused by TMAH have been reported (Chang et al. 2008). While this wastewater is toxic to humans, its nitrogen content can be a nutrient source to the water environment. As such, the Taiwan Environmental Protection Administration has established a new effluent standard of  $\text{NH}_4^+ - \text{N} < 10 \text{ mg/L}$  in water protection zones and  $< 20 \text{ mg/L}$  in other areas as applied to semiconductor and optoelectronic industries. In addition, these

industries are currently studying technologies to recover and remove TMAH from wastewater since eco-toxicity tests will be commonly required for their effluent. Therefore, TMAH-containing wastewater has to be properly treated before it can be discharged to the environment. Research conducted to treat the wastewater includes catalytic oxidation (Hirano et al. 2001), biodegradation by aerobic process (Tsuzaki and Nakamoto 1990), anaerobic process (Chang et al. 2008), aerobic and anoxic/oxic sequencing batch reactors (Lei et al. 2010), and reclamation and reuse by a combination of electro dialysis and ion exchange technology (Sugawara et al. 2002; Shibata et al. 2006).

Among existing conventional methods for treating industrial wastewater, adsorption is the most widely used process due to its simplicity and capability to produce effluents containing very low levels of dissolved organic compounds (Wibowo et al. 2007). Along with other commonly used adsorbents, such as alumina, silica gel, and zeolite, activated carbon is best known for its large surface area from internal microporosity (Do 1998). Despite the versatile use of activated carbon, currently there is still no study about TMAH adsorption on activated carbon. The only study about adsorption of TMAH uses cubic mesoporous silicate (CMS) and modified CMSs as adsorbents (Kelleher et al. 2001, 2002). The main objective of this study is to investigate the adsorption behavior of TMAH onto activated carbon.

## Materials and Methods

Different commercial activated carbons with similar surface area were obtained from Acros activated charcoal (SAC), Norit SA-2, decolorizing), Wako activated charcoal (WAC), Sigma-Aldrich activated carbon (ZAC), and Merck activated carbon (MAC). Prior to use, all activated carbons were washed with ultra pure water repetitively until the pH of the washing water was constant and the conductivity was below  $5 \mu\text{S/cm}$ . The carbons were then dried in an oven at  $105^\circ\text{C}$  overnight and sieved until the particles were below 0.149 mm (100 mesh). A stock solution was prepared by

<sup>1</sup>Graduate Student, Dept. of Chemical Engineering, National Taiwan Univ. of Science and Technology, 43 Keelung Rd., Section 4, Taipei 106, Taiwan.

<sup>2</sup>Professor, Dept. of Chemical Engineering, National Taiwan Univ. of Science and Technology, 43 Keelung Rd., Section 4, Taipei 106, Taiwan (corresponding author). E-mail: liu1958@mail.ntust.edu.tw

<sup>3</sup>Associate Professor, Dept. of Chemical Engineering, Widya Mandala Catholic Univ., Jl. Kalijudan 37, Surabaya 60114, Indonesia.

<sup>4</sup>Associate Professor, Dept. of Chemical Engineering, National Taiwan Univ. of Science and Technology, 43 Keelung Rd., Section 4, Taipei 106, Taiwan.

Note. This manuscript was submitted on September 28, 2010; approved on August 31, 2011; published online on February 15, 2012. Discussion period open until August 1, 2012; separate discussions must be submitted for individual papers. This paper is part of the *Journal of Environmental Engineering*, Vol. 138, No. 3, March 1, 2012. ©ASCE, ISSN 0733-9372/2012/3-232-238/\$25.00.

diluting aqueous solution of TMAH (25%, w/w, Megaunion) using ultra pure water (Millipore). BET surface area ( $S_{\text{BET}}$ ); micropore ( $V_{\text{mi}}$ ), mesopore ( $V_{\text{me}}$ ), and total pore volume ( $V_{\text{total}}$ ); as well as pore size distribution by density functional theory (DFT) and Barret-Joyner-Halenda (BJH) models were determined from  $N_2$  adsorption-desorption isotherms at  $-196^\circ\text{C}$  by a Quantachrome, Autosorb-1 instrument. Prior to the measurements, the activated carbons were degassed at  $200^\circ\text{C}$  in a vacuum condition for at least 24 h.  $S_{\text{BET}}$  was then determined by standard BET equation applied to the adsorption-desorption data in the relative pressure ( $P/P_0$ ) range of 0.05 to 0.3 while pore size distribution was modeled from the data in the  $P/P_0$  range of 0.05 to 0.98 using Quantachrome, Autosorb-1 software package. Isoelectric point ( $\text{pH}_{\text{IEP}}$ ) and point of zero charge ( $\text{pH}_{\text{PZC}}$ ) of the activated carbons were both determined by zeta potential measurement and pH drift method (El-Sayed and Badosz 2004), respectively, with 0.01 N NaCl (Acros) as background electrolyte.  $\text{pH}_{\text{IEP}}$  could be deduced when pH versus zeta potential line crossed the line at which zeta potential equals the zero.  $\text{pH}_{\text{PZC}}$  could be found when the final pH equals the initial pH. The widely used Boehm titration was used to semi quantitatively address surface acidity and basicity of the activated carbons (Boehm 1994, 2002).

Initial Concentrations of 20, 50, 100, 150, 200, 250, 300, and 500 mg/L were added to 50 mL of TMAH in several. Into several Erlenmeyer flasks 50 mL of TMAH solution. The pH of the solution was adjusted by 0.05 N HCl (Acros) and 0.05 N NaOH (Fischer Scientific) and 0.2 g ( $\pm 1\%$ ) of activated carbon was added into each flask then the flask was covered tightly. The flasks were shaken at  $25^\circ\text{C}$  and 80 rpm shaking speed until equilibrium was reached (24 h) in a temperature-controlled water bath shaker. Amount of TMAH adsorbent,  $q_e$  (mg/g), was then defined as

$$q_e = \frac{(C_0 - C_e) \cdot V}{m}$$

Where  $C_0$  is the initial concentration of adsorbate (mg/L),  $C_e$  is the final volume of liquid (L), and  $m$  is the mass of adsorbent (g). To ensure data validity, experiments were carried out in triplicate analysis with relative standard deviation lower than 10% in every adsorption isotherm experiment. Measurement of TMAH concentration was by ion chromatography system (Dionex ICS-1000) equipped with autosampler (Spectra System, AS-1000) with the retention time of 20 min. To study effect of temperature on adsorption, the flasks were shaken at 25, 40, and  $55^\circ\text{C}$ , respectively, with identical procedures as aforementioned.

Five hundredths mL of TMAH solution with initial concentration of 300 mg/L was put into a glass reactor. The solution in the reactor was stirred by a magnetic stirrer (Corning) at the speed of 360 rpm. The pH of the solution was adjusted by 0.05 N HCl and 0.05 N NaOH so that desired pH could be obtained. Two grams ( $\pm 1\%$ ) of activated carbon was added into the reactor and then the reactor was covered. Right after the activated carbon was added, the time was recorded. It was sampled at different intervals for analysis. The amount of TMAH adsorbed onto activated carbon at a given time,  $q_t$  (mg/g), was then defined as:

$$q_t = \frac{(C_0 - C_t) \cdot V}{m}$$

Where  $C_0$  is the initial concentration of adsorbate (mg/L),  $C_t$  is the concentration of adsorbate at a given reaction time (mg/L),  $V$  is the volume of liquid (L), and  $m$  is the mass of adsorbent (g).

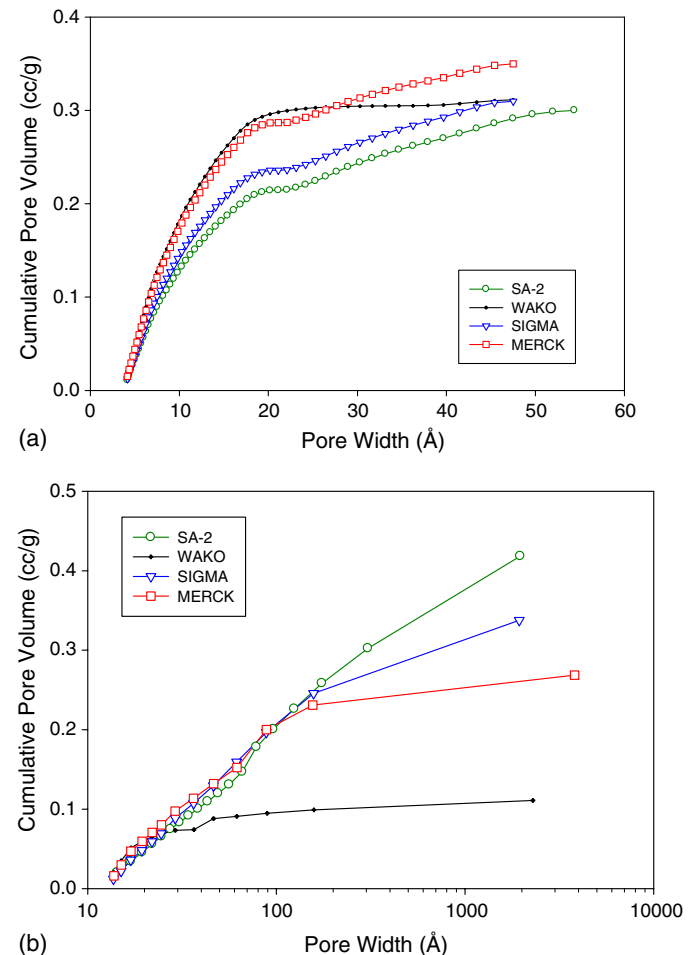
For desorption experiments, 50 mL of TMAH solution with initial concentrations of 50 mg/L was added into each Erlenmeyer flasks. The pH of the solution was controlled by 0.05 N HCl

and 0.05 N NaOH so that equilibrium pH was  $11.8 \pm 0.2$ . Fixed amount (0.2 g) of activated carbon was added into each flask and tightly covered. The flasks were shaken at 80 rpm shaking speed at  $25^\circ\text{C}$  until equilibrium was reached. After equilibrium was reached, the activated carbon was separated from the solution by filtration. It was mixed with either 50 mL of 0.1 N HCl or warm water ( $65^\circ\text{C}$ ) and was then shaken at 80 rpm for 6 h. The amount of desorbed TMAH was analyzed after samples were diluted and neutralized. The adsorbent was collected again by filtration and oven-dried at  $105^\circ\text{C}$  overnight before undergoing subsequent adsorption experiment. The cycle of such adsorption and desorption was repeated for five times in triplicate analysis.

## Results and Discussion

### Characteristics of Activated Carbons

The pore size distribution of the activated carbons was assessed by both the density function theory (DFT) and the Barret-Joyner-Halenda (BJH) models. DFT model covers pore size distribution from micropore to mesopore range (i.e., 4.2 to  $50 \text{ \AA}$ ). Fig. 1(a) illustrates that up to  $20 \text{ \AA}$ , WAC activated carbon had the highest cumulative pore volume, and there was no further increase for pore size ranging from 20 to  $50 \text{ \AA}$ . SAC, ZAC, and MAC had less cumulative pore volume for pore width  $< 20 \text{ \AA}$ , yet they showed



**Fig. 1.** (a) cumulative pore size distribution of activated carbons by DFT model; (b) cumulative pore size distribution of activated carbons by BJH model

further increase in cumulative pore volume in the pore size range of 20 to 50 Å. Similar tendency in pore size distribution of the activated carbons was also found by BJH model [Fig. 1(b).] For pore width < 20 Å, WAC had the highest pore volume, followed by MAC, ZAC, and SAC. However, as the pore size range shifted to mesopore range (20 to 500 Å), SAC had higher mesopore volume, and eventually reached the highest cumulative macropore volume (500 Å to > 2,000 Å). While for WAC, no further cumulative pore volume in mesopore and macropore range could be observed. Even though the surface areas were approximately equal, the pore size distribution of each activated carbon was different (Table 1). WAC had the lowest pore volume, followed by both MAC and ZAC, and then SAC. The ratio of mesopore to micropore revealed that WAC had the lowest value, implying that this activated carbon was more microporous as compared with other carbons. From  $pH_{PZC}$  values (Table 2), it was known that SAC was a basic carbon while the others were neutral carbons. Activated carbons with a predominantly basic surface may show an isoelectric point in the acidic range since electrokinetic behavior is determined by the charge on the external surface of the carbon particles (Corapcioglu and Huang 1987; Boehm 2002). Thus, as depicted in Table 2, the  $pH_{IEP}$  of activated carbons was always lower than  $pH_{PZC}$ . The nature of oxygen functional groups distribution caused denser oxygen surface functional groups on the external surface (Menéndez et al. 1995). Therefore, the difference between  $pH_{PZC}$  and  $pH_{IEP}$  can be regarded as a rough measure of the surface charge distribution of carbon. Positive values indicate more negatively charged external particle surfaces than internal ones and values close to zero imply a more homogenous distribution of the surface charges (Menéndez et al. 1995). Furthermore, Boehm titration results showed good agreement with the trend of  $pH_{PZC}$ , which were in the order of MAC < WAC < ZAC < SAC. The acidity of SAC and WAC was predominantly contributed by carboxyliclike group, while ZAC and MAC by phenoliclike group. Even though WAC activated carbon had the densest carboxyliclike group,  $pH_{PZC}$  of MAC (i.e., 6.4) was lower than WAC (i.e., 7.4) since MAC had the lowest amount of basic functional group. In contrast, SAC possessed highest basic functional group despite its high carboxyliclike group content, making it a basic activated carbon. The amount of acidic and basic surface functional

**Table 1.** Specific Surface Area and Pore Volume Distribution of the Activated Carbons

Activated carbon	$S_{BET}$ m <sup>2</sup> /g	Pore volume (cc/g)			
		$V_{mi}$ (DFT-Method)	$V_{me}$ (BJH-method)	$V_{total}$	$V_{me}/V_{mi}$
SAC	619.80	0.1487	0.4185	0.5713	2.81
WAC	862.19	0.2544	0.1110	0.3609	0.44
ZAC	670.09	0.1698	0.3375	0.5119	1.99
MAC	854.94	0.1945	0.2686	0.4887	1.38

**Table 2.** Surface Properties and Functional Groups of the Activated Carbons

$pH_{PZC}$	$pH_{IEP}$	Acidic groups (meq/g)				Basic group (meq/g)
		Carboxylic	Lactonic	Phenolic	Total	
10.1	4.6	0.124	0.070	0.038	0.232	1.058
7.4	5.3	0.179	0.134	0.160	0.473	0.459
7.9	7.6	0.011	0.110	0.416	0.537	0.652
6.4	3.6	0.005	0.026	0.434	0.465	0.077

groups affected the total net surface charge of activated carbons significantly.

### Kinetic Study

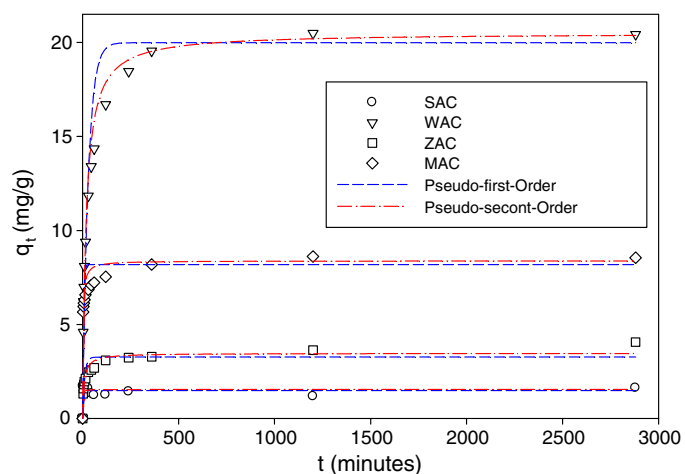
One of the widely used models to describe adsorption kinetics is pseudo-first-order model. The equation has the form after integration (Ho and McKay 1998)

$$q_t = q_e(1 - \exp(-k_1t)) \quad (1)$$

where  $k_1$  is the rate constant for the first order ( $\text{min}^{-1}$ )  $q_e$  is equilibrium adsorption capacity, and  $q_t$  is measured adsorption capacity at any time. Another is the pseudo-second-order kinetic model is as follows (Ho 2006):

$$q_t = q_e \left( \frac{q_e k_2 t}{1 + q_e k_2 t} \right) \quad (2)$$

TMAH adsorption on SAC, WAC, ZAC, and MAC with respect to reaction time is illustrated and fitted by pseudo-first-order and pseudo-second-order kinetic models (Fig. 2). It can be found that pseudo-second-order kinetic model fitted better than pseudo-first-order for all kinds of activated carbons, with correlation coefficients higher than 0.90 (Table 3). The models were not applicable for SAC since the type of kinetic curves was pseudo rectangular; indicating that the adsorption rapidly approached equilibrium (Chen and Wu 2004). Similar behavior was also observed on adsorption kinetics of TMAH on CMSs (Kelleher et al. 2001), cationic surfactant



**Fig. 2.** Pseudo-first-order and pseudo-second-order kinetic plot of TMAH adsorption on SAC, WAC, ZAC, and MAC ( $C^0 = 300 \pm 30$  mg/L, volume = 500 mL, adsorbent dose = 4 g/L, pH =  $11.8 \pm 0.2$ )

**Table 3.** Reaction Kinetics of TMAH Adsorption on SAC, WAC, ZAC, and MAC

Activated carbon	Pseudo-first-order $q_t = q_e(1 - \exp(-k_1t))$			Pseudo-second-order $q_t = q_e \left( \frac{q_e k_2 t}{1 + q_e k_2 t} \right)$		
	$q_e$ (mg/g)	$k_1$ (1/min)	$R^2$	$q_e$ (meq/g)	$k_2$ (meq/g)	$R^2$
SAC	1.48	N/A	N/A	1.54	N/A	N/A
WAC	19.99	0.0344	0.8850	20.50	0.0027	0.9719
ZAC	3.26	0.0945	0.8158	3.46	0.0367	0.9214
MAC	8.19	0.2435	0.8021	8.38	0.0496	0.9125

(i.e., cetyltrimethylammonium bromide) on activated carbon (Yalçın et al. 2005), as well as ammonium on sawdust (Wahab et al. 2010). In the current study, it is because generally experimental data are better represented by pseudo-second-order rather than pseudo-first-order model has been well discussed (Plazinski et al. 2009; Kosasih et al. 2010).

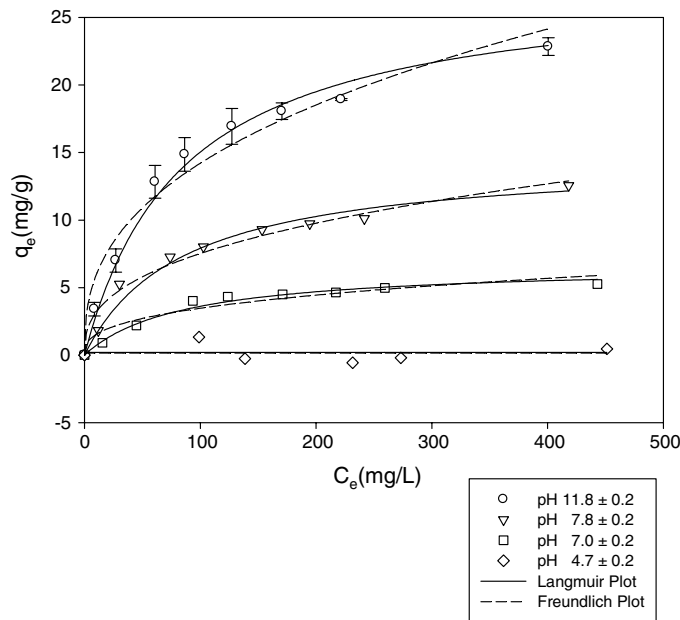
### Adsorption Isotherm

Adsorption isotherms of WAC at different equilibrium pH are shown in Fig. 3. The experimental data were fitted by two-parameter isotherms, i.e., Langmuir and Freundlich equations. The Freundlich equation is written as follows:

$$q_e = K_f \cdot C_e^{1/n} \quad (3)$$

where  $q_e$  is the mass of substance adsorbed per unit mass of adsorbent,  $C_e$  is the equilibrium concentration of the solute being adsorbed, while  $K_f$  and  $n$  are empirical constants. The Langmuir isotherm is in the form of

$$q_e = q_0 \cdot \frac{b \cdot C_e}{1 + b \cdot C_e} \quad (4)$$



**Fig. 3.** Adsorption isotherms of TMAH on WAC at different pH as fitted by Langmuir and Freundlich isotherms (adsorbent dose = 4 g/L, reaction time = 24 hours)

where  $C_e$  is the equilibrium concentration (mg/L),  $q_e$  is the amount adsorbed at equilibrium (mg/g), and  $q_0$  and  $b$  are Langmuir constants related to the adsorption capacity and energy of adsorption, respectively. Langmuir and Freundlich fitting parameters are summarized in Table 4. The pH greatly influenced TMAH adsorption so that the higher the pH, the higher the TMAH adsorption capacity of WAC activated carbon. At equilibrium pH of  $4.7 \pm 0.2$ , no adsorption of TMAH occurred, while at equilibrium pH of  $11.8 \pm 0.2$ , the adsorption capacity was found to be the highest. The Langmuir model fitted very well with every adsorption isotherm data of WAC, with correlation coefficient value higher than 0.97, better than Freundlich equation ( $0.91 < R^2 < 0.97$ ). Adsorption density parameter,  $q_0$ , changed with respect to pH, while the adsorption energy constant,  $b$ , was constant within entire pH range studied. Aside from fitting capability, the  $n$  values of the adsorption were between 1 and 10, indicating favorable adsorption isotherm type. A suitable fitting by Langmuir isotherm indicated formation of monolayer coverage of TMAH molecules at the adsorption sites of the adsorbent (Do 1998). Similar good correlation with Langmuir isotherm was also found by Kelleher et al. (2001, 2002) on their CMSs in the range of equilibrium concentration of about 0 to ca. 300 mg/L of TMAH. Compared with the cubic mesoporous silicates (CMSs), the adsorption capacities of WAC activated carbon were less than one tenth. CMSs surface-area-to-volume ratios were very large, which suggests the dominant presence of microporosity, and the ratios were much higher than WAC. This fact indicates that microporosity had a very important role in adsorption of TMAH.

### Effect of Temperature

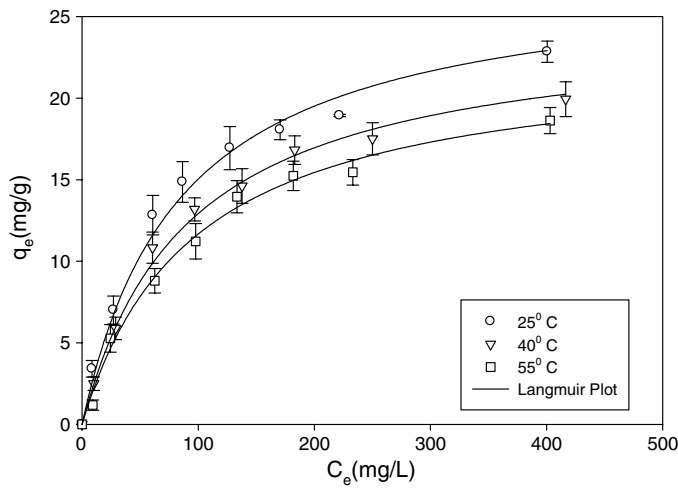
As depicted in Fig. 4 and Table 5, Langmuir model fitted well adsorption of TMAH on WAC activated carbon at different temperature. The increase of temperature from 25°C to 55°C was disadvantageous not only to the  $b$  value, but also to the maximum adsorption capacity. The decrease in maximum adsorption capacity with the increase of temperature could be due to increasing tendency of the TMAH molecules to escape from the solid phase to the bulk phase as adsorbate molecules possessed higher total energy (Ho and McKay 1998) and that the adsorption process was dominated by physisorption, resulting in the bonding between the adsorbate and the active adsorbent sites that weakens at higher temperature, in contrast with chemisorption (Kosasih et al. 2010). Thermodynamic analysis of the adsorption reaction can be assessed utilizing van't Hoff equation:

$$\ln b = -\frac{\Delta H^\circ}{RT} + \frac{\Delta S^\circ}{R} \quad (5)$$

where  $b$  is the Langmuir constant,  $\Delta H^\circ$  the standard enthalpy change,  $\Delta S^\circ$  the standard entropy change,  $R$  the gas constant,

**Table 4.** Isotherm Parameters of TMAH Adsorption on WAC at Different pH

Isotherms	Parameters	pH			
		11.8 ± 0.2	7.8 ± 0.2	7.0 ± 0.2	4.7 ± 0.2
Freundlich ( $q_e = K_f \cdot C_e^{1/n}$ )	$K_f$ (mg/g)(L/mg) <sup>1/n</sup>	2.44	1.33	0.67	N/A
	$n$	2.61	2.65	2.81	N/A
	$1/n$	0.383	0.377	0.356	N/A
	$R^2$	0.9652	0.9687	0.9143	N/A
Langmuir ( $q_e = q_0 \cdot \frac{b \cdot C_e}{1 + b \cdot C_e}$ )	$q_0$ (mg/g)	27.77	14.62	6.68	N/A
	$b$ (L/mg)	0.0118	0.0118	0.0118	N/A
	$R^2$	0.9902	0.9768	0.9803	N/A



**Fig. 4.** Effect of temperature on adsorption isotherms of WAC (adsorbent dose = 4 g/L, pH = 11.8 ± 0.2, reaction time = 24 hours)

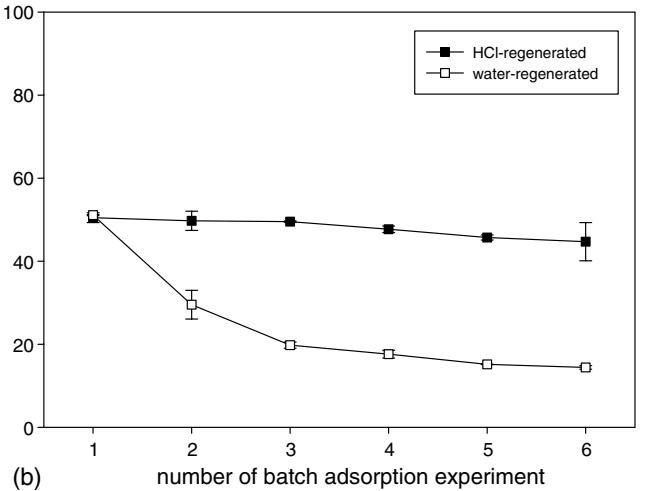
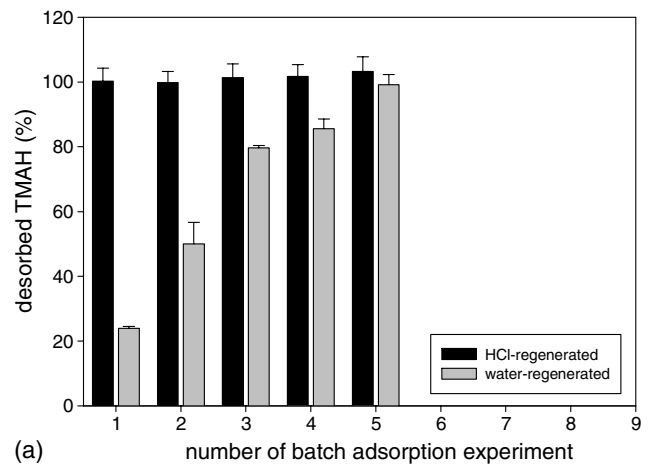
**Table 5.** Langmuir Constant of TMAH Adsorption at Different Temperature

T (°C)	$q_0$ (mg/g)	$b$ (L/mg)	$R^2$
25	27.77	0.0118	0.9902
40	23.53	0.0110	0.9956
55	22.81	0.0105	0.926

and  $T$  the temperature. By the slope and intercept of linear regression of van't Hoff equation, the thermodynamic parameters in Table 6 could be determined. It is clear that the standard Gibbs' free energy change were all negative along the temperature variations, indicating a spontaneous nature of TMAH adsorption on WAC activated carbon (Do 1998). The negative value of the enthalpy change revealed that the adsorption was exothermic in nature. On the contrary, CMSs showed poor adsorption when the temperature was set to low (i.e., 7°C) and when the temperature was increased from 23 to 40°C,  $b$  and  $q_0$  were increased, suggesting endothermic process (Kelleher et al. 2001). The typical enthalpy change in physisorption is in the range of 5–40 kJ/mol (1.19–9.56 kcal/mol) and typical enthalpy change in ion exchange is below 2 kcal/mol (Helfferich 1995; Inglezakis and Pouloupoulos 2006). The adsorption enthalpy range was in the range of a combination of physisorption and ion exchange. Positive value of  $\Delta S^\circ$  suggests that adsorption process exhibits randomness behavior at the solid/solution surface (Kelleher et al. 2001; Kosasih et al. 2010).

### Desorption Study

In five cycles of adsorption-desorption experiments, the TMAH-saturated activated carbon could be easily regenerated by 0.1 N HCl, while the regeneration could not be completed by warm water [Figs. 5(a) and 5(b)]. By using 0.1 N HCl, the amount of desorbed



**Fig. 5.** (a) regeneration of WAC by 0.1 N HCl and water ( $C^\circ = 50 \pm 5$  mg/L, adsorbent dose = 4 g/L, pH = 11.8 ± 0.2, reaction time = 24 hours); (b) batch adsorption cycles of WAC ( $C^\circ = 50 \pm 5$  mg/L, adsorbent dose = 4 g/L, adsorbent dose = 4 g/L, reaction time = 24 hours)

TMAH could always reach 100% even after five cycles of adsorption-desorption. Even though there was an increase in the percentage of desorbed TMAH with the increase of cycle time, regeneration efficiency of the water actually stayed the same while WAC activated carbon was actually more and more exhausted. Eventually, removal efficiency of TMAH stabilized at a low value (ca. 15%), with the same amount of TMAH that could be desorbed. Overall, this relatively easy desorption suggested that the bond between TMAH and activated carbon surface was a weak physical one and the adsorption was reversible.

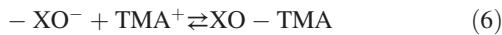
### Mechanism

Since activated carbon is amphoteric in nature, its surface charge changes as the solution pH varies. Since TMAH remains

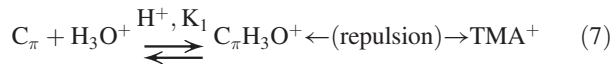
**Table 6.** Thermodynamic Parameters of TMAH Adsorption at Different Temperature

T (°C)	$1/T(1/K)$	$b'$ L/mol	$\ln b'$	$\Delta G^\circ$ (kcal/mol)	$\Delta H^\circ$ (kcal/mol)	$\Delta S^\circ$ (cal/[mol.k])
25	0.003354	1075.57	6.9806	-4.13	-0.76	11.32
40	0.003193	1002.65	6.9104	-4.30		
55	0.003047	957.08	6.8639	-4.47		

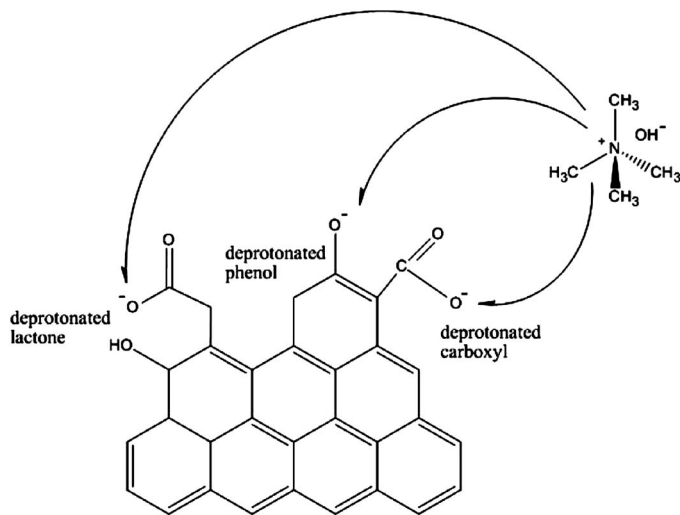
permanently ionized over any pH range, its adsorption depended strongly on the pH of the solution that determines the surface charge of the activated carbons. When the pH was shifted to a region lower than the  $pH_{PZC}$  of the WAC activated carbon, the adsorption capacity decreased significantly until there was practically no adsorption. Since no adsorption of TMAH could be observed when the pH was lower than  $pH_{PZC}$  (i.e., at  $pH = 4.7 \pm 0.2$ ), it suggested that hydrophobic interaction was not a dominant driving force for the adsorption of TMAH on activated carbon. On the contrary, as the pH shifted from  $pH_{PZC}$  to higher pH region, the adsorption density increased dramatically. Effect of pH on TMAH adsorption suggested that the electrostatic interaction was the main driving force for the adsorption. As demonstrated by results from Boehm titration, the activated carbons contained polar functional groups, such as carboxylic, lactonic, and phenoliclike groups. Therefore, the reaction between TMAH and the surface of the activated carbons can be represented as



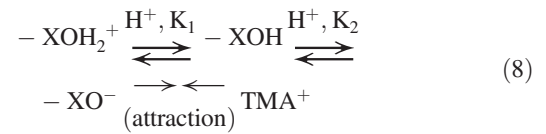
$X$  could be carbon or oxygen element, representing carboxylic, lactonic, and phenoliclike functional groups whose amounts were semi quantitatively characterized by Boehm titration. For adsorption at high pH, complete ionization of the carboxylic, lactonic, and phenolic functional groups on the surface of activated carbon unleashed its maximum surface charge density. This situation is schematically illustrated at Fig. 6. In general, adsorption of TMAH on activated carbon can be visualized as follow. First, delocalized  $\pi$  electrons of the graphite planes may act as Lewis basic sites accepting protons as governed by this reaction scheme (Lopez-Ramon et al. 1999; László 2001):



which is disadvantageous to TMAH adsorption. Secondly, the presence of acidic functional group is beneficial to TMAH adsorption, particularly when solution pH is higher than  $pH_{PZC}$ . It was the negatively charged surface oxides tend to attract  $TMAH^+$  as follows (Lopez-Ramon et al. 1999):



**Fig. 6.** Schematic representation of  $TMAH^+$  attracted to ionized acidic surface functional groups of the activated carbon



This was the first work in which activated carbon was utilized for TMAH adsorption. Judging from experimental results, activated carbon was in general very effective, particularly those with micropores and low  $pH_{PZC}$ . It can be used as a pretreatment of TMAH-containing wastewater prior to its biological treatment unit, so that the wastewater becomes more biodegradable. Or, it can be a post-treatment following biological process to further polish secondary effluent to lower the effluent eco-toxicity. Some column study is needed to assess adsorption reactions in more detail before pilot study. Subsequent treatment of loaded activated carbon is relatively easy if the removal of TMAH only is desired. Since TMAH is completely decomposed at temperature of  $135\text{--}140^\circ\text{C}$ , the TMAH-exhausted activated carbon can be heat-treated and washed instead of incinerated (Feng et al. 2008). Alternatively, since TMAH could be desorbed completely by 0.1 N HCl, the concentrated wastewater generated from activated carbon desorption can be further recovered via membrane processes. Surely, further feasibility study and cost analysis should be considered.

## Conclusions

Adsorption of TMAH on activated carbons was carried out in this study. Based on the experimental results and discussion, the following conclusions could be drawn:

1. In the adsorption of TMAH on activated carbons, the activated carbon with micropores was favored.
2. The adsorption kinetics could be represented by pseudo-second-order model.
3. The adsorption of TMAH on activated carbon followed monolayer adsorption as indicated by good fitting by Langmuir isotherm. The highest adsorption capacity among activated carbons studied was  $27.77\text{ mg/g}$  as achieved by WAC.
4. Adsorption of TMAH was noticeable only when solution pH was higher than the  $pH_{PZC}$  of the activated carbon, and it was mainly driven by electrostatic adsorption mechanism.
5. Adsorption process was spontaneous and exothermic in nature with standard enthalpy change in the typical range between physical adsorption and ion exchange.
6. From the desorption study, exhausted activated carbon could be regenerated by 0.1 N HCl without significant reduction in TMAH removal. The desorption study also confirmed that the bonding between TMAH and the active adsorption sites of activated carbon was weak and of physical nature.

## Acknowledgments

Devarly Prahas would like to acknowledge the full international scholarship awarded by National Taiwan University of Science and Technology during his master study.

## References

- Boehm, H. P. (1994). "Some aspects of the surface chemistry of carbon blacks and other carbons." *Carbon*, 32(5), 759–769.
- Boehm, H. P. (2002). "Surface oxides on carbon and their analysis: A critical assessment." *Carbon*, 40(2), 145–149.

- Chang, K. F., Yang, S. Y., You, H. S., and Pan, J. R. (2008). "Anaerobic treatment of tetra-methylammonium hydroxide (TMAH) containing wastewater." *IEEE Trans. Semicond. Manuf.*, 21(3), 486–491.
- Chen, J. P., and Wu, S. (2004). "Acid/base-treated activated carbons: Characterization of functional groups and metal adsorptive properties." *Langmuir*, 20(6), 2233–2242.
- Corapcioglu, M. O., and Huang, C. P. (1987). "The surface acidity and characterization of some commercial activated carbons." *Carbon*, 25(4), 569–578.
- Do, D. D. (1998). *Adsorption Analysis: Equilibria and Kinetics*, Imperial College Press, London.
- El-Sayed, Y., and Bandosz, T. J. (2004). "Adsorption of valeric acid from aqueous solution onto activated carbons: Role of surface basic sites." *J. Colloid Interface Sci.*, 273(1), 64–72.
- Feng, H., Huang, C., and Xu, T. (2008). "Production of tetramethyl ammonium hydroxide using bipolar membrane electrodialysis." *Ind. Eng. Chem. Res.*, 47(20), 7552–7557.
- Helfferich, F. (1995). *Ion exchange*, Dover Publications, NY, 166.
- Hirano, K., Okamura, J., Taira, T., Sano, K., Toyoda, A., and Ikeda, M. (2001). "An efficient treatment technique for TMAH wastewater by catalytic oxidation." *IEEE Trans. Semicond. Manuf.*, 14(3), 202–206.
- Ho, Y. (2006). "Review of second-order models for adsorption systems." *J. Hazard. Mater.*, 136(3), 681–689.
- Ho, Y. S., and McKay, G. (1998). "Kinetic models for the sorption of dye from aqueous solution by wood." *Process Saf. Environ. Prot.*, 76(2), 183–191.
- Inglezakis, V. J., and Pouloupoulos, S. G. (2006). *Adsorption, ion exchange, and catalysis: Design operations and environmental applications*, Elsevier, Amsterdam, 31–33.
- Kelleher, B. P., Doyle, A. M., Hodnett, B. K., and O'Dwyer, T. F. (2002). "An investigation into the adsorption characteristics of grafted mesoporous silicates for the removal of tetramethyl ammonium hydroxide from aqueous solution." *Adsorpt. Sci. Technol.*, 20(8), 787–796.
- Kelleher, B. P., Doyle, A. M., O'Dwyer, T. F., and Hodnett, B. K. (2001). "Preparation and use of a mesoporous silicate material for the removal of tetramethyl ammonium hydroxide (TMAH) from aqueous solution." *J. Chem. Technol. Biotechnol.*, 76(12), 1216–1222.
- Kosasih, A. N., Febrianto, J., Sunarso, J., Ju, Y. H., Indraswati, N., and Ismadji, S. (2010). "Sequestering of Cu(II) from aqueous solution using cassava peel (*Manihot esculenta*)." *J. Hazard. Mater.*, 180(1–3), 366–374.
- László, K., Tombácz, E., and Josepovits, K. (2001). "Effect of activation on the surface chemistry of carbons from polymer precursors source." *Carbon*, 39(8), 1217–1228.
- Lei, C. N., Whang, L. M., and Chen, P. C. (2010). "Biological treatment of thin-film transistor liquid crystal display (TFT-LCD) wastewater using aerobic and anoxic/oxic sequencing batch reactors." *Chemosphere*, 81(1), 57–64.
- Lopez-Ramon, M. V., Stoeckli, F., Moreno-Castilla, C., and Marin, F. (1999). "On the characterization of acidic and basic surface sites on carbons by various techniques." *Carbon*, 37(8), 1215–1221.
- Menéndez, J. A., Illan-Gomez, M. J., León y León, C. A., and Radovic, L. R. (1995). "On the difference between the isoelectric point and the point of zero charge of carbons." *Carbon*, 33(11), 1655–1659.
- Plazinski, W., Rudzinski, W., and Plazinska, A. (2009). "Theoretical models of sorption kinetics including a surface reaction mechanism: A review." *Adv. Colloid Interface Sci.*, 152(1–2), 2–13.
- Shibata, J., Murayama, N., and Matsumoto, S. (2006). "Recovery of tetramethyl ammonium hydroxide from waste solution by ion exchange resin." *Resour. Process.*, 53(4), 199–203.
- Sugawara, H., Tajima, Y., and Ohmi, T. (2002). "A study on reclaimed photoresist developer using an electrodialysis method." *Jpn. J. Appl. Phys.*, 41(Part 1, No. 4B), 2374–2379.
- Thong, J. T. L., Choi, W. K., and Chong, C. W. (1997). "TMAH etching of silicon and the interaction of etching parameters." *Sens. Actuators, A*, 63(3), 243–249.
- Tsuzaki, M., and Nakamoto, S. (1990). "Biodegradation of tetramethylammonium, an ingredient of a positive photoresist developer, by bacteria isolated from activated sludge." *NEC Res. Dev.*, (97), 8–12.
- Wahab, M. A., Jellali, S., and Jedidi, N. (2010). "Ammonium biosorption onto sawdust: FTIR analysis, kinetics and adsorption isotherms modeling." *Bioresour. Technol.*, 101(14), 5070–5075.
- Wibowo, N., Setyadhi, L., Wibowo, D., Setiawan, J., and Ismadji, S. (2007). "Adsorption of benzene and toluene from aqueous solutions onto activated carbon and its acid and heat treated forms: Influence of surface chemistry on adsorption." *J. Hazard. Mater.*, 146(1–2), 237–242.
- Yalçin, M., Gürses, A., Dogar, C., and Sözbilir, M. (2005). "The adsorption kinetics of cetyltrimethylammonium bromide (CTAB) onto powdered active carbon." *Adsorption*, 10(4), 339–348.