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**10 Solubilities of 3-acetylpyridine in supercritical carbon dioxide at several temperatures and pressures: Experimental and modeling**

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3-Acetylpyridine Solubility Density-based model Equations of state Phase equilibria abstract 3-Acetylpyridine (methyl 3-pyridyl ketone) is one of the important compounds to impart flavor and fragrance in various food products. In this work, the solubility data of 3-acetylpyridine in supercritical carbon dioxide (SC-CO<sub>2</sub>) were experimentally measured at several temperatures

**19(313.15 K, 323.15 K, 333.15 K, and 343.15 K) and pressures from 10 MPa to 26 MPa** under static mode. **The experimental**

solubilities over the measurement range were correlated using Chrastil and Del Valle and Aguilera density-dependent models while phase equilibrium behavior of the studied system was interpreted by Peng–Robinson equation of state (PR-EoS) incorporated with quadratic and Stryjek–Vera mixing rules. The agreement between the model predictions and experimental solubilities was assessed in respect to root-mean-square deviation (RMSD) and consistency of physical meaning of model parameters. Concerning phase equilibria of the studied supercritical system, PR-EoS incorporated with quadratic mixing rule was superior to PR-EoS incorporated with Stryjek–Vera mixing rule and capable to describe the dependency of adjustable binary interaction parameters with temperature. © 2013 Elsevier B.V. All rights reserved. 1. Introduction Pyridines represent a class of heterocyclic organic compounds that significantly contribute to the organoleptic properties of many food products. Of particular interest is 3-acetylpyridine (or methyl 3-pyridyl ketone), a colorless to yellowish clear liquid with peanut-like odor and can be naturally found in roasted hazelnuts/filberts and cocoa [1,2]. This compound is frequently used as a flavoring agent and to impart a distinctive scent in the baked goods, candies, puddings, and non-alcoholic beverages. The isolation techniques of 3-acetylpyridine from grounded roasted hazelnuts/filberts are mostly conducted by steam or molecular distillation processes [2]. Denaturation risk and low purity of the final product are two main disadvantages of those conventional processes, which drive the searching of alternative techniques to appropriately isolating this heat-labile compound. Over the past decades, supercritical fluid extraction has gained increasing awareness as a promising method for isolating valuable bioactive compounds from various natural products for \*

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3@yahoo.com (S. Ismadji). 0378-3812/\$ – see front matter © 2013 Elsevier B.V.  
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2013.06.036 pharmaceutical and food industries. This technique is considered greener compared to conventional extraction processes which usually involve the use of harmful organic solvents and fairly high operating temperature [3]. Another uniqueness of supercritical fluid extraction is that this process generates zero waste because the solvent is completely recyclable and can be conducted at nearly-ambient temperature, thus ensuring the stability and quality of the final product. To date, carbon dioxide (CO<sub>2</sub>) is the most widely used supercritical solvent due to its practical advantages such as non-toxic, non-flammable, and non-explosive, environmentally acceptable, cheap, low critical temperature (T<sub>c</sub> of 304.2 K), and easily obtained in ultra high purity [3,4]. The information about solubilities of the compounds of interest in supercritical solvents at wide range of pressures and temperatures is of crucial importance for the process design and optimizing supercritical-based processes, particularly for industrial practice.

4To the best of our knowledge, there is no previous report available in the literatures about the solubilities of

3-acetylpyridine in supercritical carbon dioxide. In the present work, the solubilities of 3-acetylpyridine in supercritical carbon dioxide is measured in a static operation mode at four temperatures ranging

2from 313.15 K to 343.15 K and several pressures range between 10 MPa and 26 MPa. The

experimental solubility data were modeled with Chrastil and Del Valle and Aguilera semi-empirical density-dependent correlations and Peng–Robinson cubic equation of state Table 1 Some chemical information about 3-acetylpyridine compound. V-2 E F Name 3-Acetylpyridine CAS number 350-03-8 Purity ≥98%, food grade C Physical appearance Clear, dark yellow liquid Organoleptic Nutty; sweet B Empirical formula C<sub>7</sub>H<sub>7</sub>NO Molecular mass 121.14 g/mol Boiling point 493.15 K A D incorporated with quadratic and Stryjek–Vera mixing rules was V-1 applied to interpret phase equilibria of the studied system. 2. Materials and methods 2.1. Chemicals Food grade 3-acetylpyridine was purchased from Sigma–Aldrich chemical company, Singapore without further purification. Some information about this flavoring compound are given in Table 1. Food grade carbon dioxide (99.98% pure) was supplied by a local gas company in a deep tube liquid supply cylinder. Ethanol was

21Fig. 1. Schematic of experimental set-up for solubility

measurement in high pres- obtained from Sigma–Aldrich with a purity of 96% and used as a sure condition (dip tube liquid CO<sub>2</sub> supply cylinder (A); high pressure liquid liquid medium for trapping 3-acetylpyridine. metering pump (B); high pressure equilibration cylinder (C); pressure transducer (D); thermostated oven (E); receiving vial (F); valve – (V-1) inlet line, (V-2) outlet 2.2. Solubility measurements line). The schematic of experimental set-up for static measurements on the receiving vial to allow the expanded gas to exit and total of 3-acetylpyridine solubilities in supercritical carbon dioxide is volume of solute-free ambient gas was measured with a calibrated displayed in Fig. 1. The supercritical apparatus consists of a vac- wet gas flow meter (±0.05 L) at a known pressure and tempera- uum pump (GAST DOA-P504-BN), a calibrated wet gas

flow meter ture. Care was taken to prevent sample precipitation during the (ZEAL DM3B), a high pressure liquid metering pump with two recip- decompression of high-pressure stream by gently heating the dis- roccating pistons (Eldex AA-100-S-2), a pressure transducer (Druck charge line and flushing with fresh CO<sub>2</sub>). The ethanol containing PTX 611) connected to digital process indicator (Druck DPI 280), 3-acetylpyridine was then analyzed on a Shimadzu GC-2014 gas a thermostated oven (Mettler UM400), and a 150 ml high pres- chromatograph under prescribed condition. sure double-ended sample cylinder (17.22 cm long × 4.85 cm i.d.). The quantification of 3-acetylpyridine concentration in ethanol All fittings and tubings are made of 316SS-grade stainless steel was conducted by means of gas chromatography analysis, adopting (Swagelok, USA). The maximum working temperature and work- operation condition reported by Kulshreshtha and Moldoveanu [5]. ing pressure of the system is 423.15 K and 40 MPa, respectively. The The gas chromatograph was equipped with a split/splitless injector, measurement uncertainties were ±0.01 MPa for pressure and ±1 K an Agilent narrowbore

18DB-Wax capillary column (30 m × 0.25 mm; for temperature. 0.25

?m) as the stationary phase, and a flame ionization detector Briefly,

2a known amount of cotton (±1 g) was wetted with

liq- (FID). The

16injection volume was 1 ?l in splitless mode. The

injector uid

23-acetylpyridine (±1 ml) and packed into the sample cylinder.

and FID temperature was operated at 280 °C and 300 °C, respec- Both ends of the cylinder were fitted with sintered stainless steel tively. Highly pure helium (99.98%) was employed as the carrier gas filter discs (50 ? m) to prevent any entrainment loss. Then, the sys- at pressure of 134.9 kPa and a column flow of 2 ml/min. The initial tem was evacuated with a vacuum pump. The evacuated system column oven temperature was set at 40 °C for 5 min, then ramped was heated to predetermined temperatures. Subsequently, liquid to 110 °C with a heating rate of 4 °C/min and increased to 230 °C at CO<sub>2</sub> was compressed and delivered passing through the tubing to

1610 °C/min and held for 5.5 min. Total program time was 40 min.

the sample cylinder

18at a constant flow rate of 10 ml/min (mea- Identification of

3-acetylpyridine peak in the samples was con- sured at 6.8 MPa and 303.15 K) until desired pressures reached. The ducted by matching retention time with that previously obtained system pressure was monitored real-time and the system tempera- by the injection of authentic standard. Calibration curve was

pre- ture was controlled by a built-in PID-type temperature controller of pared by injecting five standard solutions of 3-acetylpyridine with the oven to maintain isobar and isothermal conditions. Preliminary different concentrations (0.1–10 mg/L) and plotting the average experiments show that equilibrium was established within 2 h. The peak areas as ordinate vs. the concentrations of each analyte as equilibrated supercritical phase was thereafter released from the abscissa. Then, the concentration of 3-acetylpyridine in the sam- sample cylinder at the same condition by fresh CO<sub>2</sub> into a receiving ples (wt.%) was determined from as- prepared calibration curve vial containing some volume of 96% ethanol to completely separate based on the detected peak area. The solubilities at a given pres- 3-acetylpyridine and carbon dioxide. A vent needle was mounted sure and temperature were determined by quantifying the amount B. Agustin et al. / Fluid Phase Equilibria 354 (2013) 127–132 129 Table 2 Experimental

## 2solubilities of 3-acetylpyridine in supercritical

CO<sub>2</sub> over the measure- ment range. P (MPa) Solubility ( $\times 10^{-2}$ ) (kg/m<sup>3</sup>) 313.15 K 323.15 K 333.15 K 343.15 K 10 1.541  $\pm$  0.005 12 1.915  $\pm$  0.009 14 2.426  $\pm$  0.007 16 2.790  $\pm$  0.004 18 2.986  $\pm$  0.008 20 3.593  $\pm$  0.044 22 3.916  $\pm$  0.033 24 4.717  $\pm$  0.008 26 5.423  $\pm$  0.018 1.152  $\pm$  0.013 1.715  $\pm$  0.051 2.683  $\pm$  0.049 3.622  $\pm$  0.022 4.416  $\pm$  0.008 5.055  $\pm$  0.005 6.041  $\pm$  0.036 6.757  $\pm$  0.051 7.754  $\pm$  0.044 0.904  $\pm$  0.006 1.473  $\pm$  0.022 2.722  $\pm$  0.019 4.258  $\pm$  0.022 5.754  $\pm$  0.036 7.116  $\pm$  0.018 8.682  $\pm$  0.024 9.824  $\pm$  0.053 11.292  $\pm$  0.084 0.69  $\pm$  0.012 1.38  $\pm$  0.044 2.93  $\pm$  0.036 4.75  $\pm$  0.027 7.03  $\pm$  0.006 9.42  $\pm$  0.008 11.37  $\pm$  0.018 13.91  $\pm$  0.043 16.19  $\pm$  0.055 of 3-acetylpyridine in ethanol for a given volume of CO<sub>2</sub>. All measurements were repeated three times at identical operational conditions and presented as mean  $\pm$ SD for three replicate determi- nations. 3.

Results and discussion 3.1. Solubilities of 3-acetylpyridine in supercritical carbon dioxide The quantity of a particular compound that can dissolve in a par- ticular supercritical solvent to form a saturated supercritical phase strongly depends on the solvent density. In this regards, the den- sities of supercritical CO<sub>2</sub> at various temperatures and pressures were estimated using

## 22Stryjek–Vera modification of Peng–Robinson (PRSV) equation of state

and the results are given in Supplementary Material Table S1. The

## 22use of PRSV equation of state

is generally superior to other equations for estimating the fluid densities of many materials in moderately non-ideal systems, typically nonpo- lar ones with relatively accurate correlation results [6]. Moreover, PRSV equation of state is convenient to use since it only requires the knowledge of critical properties and acentric factor for the generalized parameters and computationally efficient to solve this equation. The critical pressure, critical temperature, and acen- tric factor for carbon dioxide and 3-acetylpyridine is 7.39 MPa, 304.25 K, 0.225 and 4.47 MPa, 729.15 K, and 0.464, respectively. At isothermal condition (e.g., 313.15 K), the increase of

## 2pressures from 10 MPa to 20 MPa causes the solvent power of the

fluid to increase from  $(1.541 \pm 0.005) \times 10^{-2}$  kg/m<sup>3</sup> to  $(3.593 \pm 0.044) \times 10^{-2}$  kg/m<sup>3</sup>. Similar trends were noticed in other three temperatures studied. The experimental solubilities of 3-acetylpyridine in SC-CO<sub>2</sub> over the measurement range were summarized in Table 2. It can be seen that the solubilities increased at elevated pressures and tempera- tures, except those at two pressures of 12 MPa and 14 MPa. It is clearly

understandable that increasing pressure at a constant temperature enhances the solvent density, leading to increased amount of solute dissolved in the solvent phase. Meanwhile, the system temperature has different effects on solubility

**14** due to the influences of temperature on the vapor pressure, the density of supercritical solvent, and the molecular interaction of supercritical phase. Near the critical region of

CO<sub>2</sub> (304.25 K, 7.39 MPa), a small temperature rise

**2** from 313.15 K to 323.15 K at 10 MPa may lead to

a large fall in the solvent density (565 kg/m<sup>3</sup> to 376 kg/m<sup>3</sup>) and a resulting change in 3-acetylpyridine solubilities ( $(1.541 \pm 0.005) \times 10^{-2}$  to  $(1.152 \pm 0.013) \times 10^{-2}$  kg/m<sup>3</sup>). Increasing temperature would increase the solute's volatility and the diffusivities of both solvent and solute species, but cause the reduction in solvent capacity. These opposing effects resulted in the crossover phenomena in the solubility isotherms as illustrated in Fig. 2. The actual solubility data and regression models are designated as dot symbols and dashed lines, respectively. A near-linear relationship between pressure and mole fraction of 3-acetylpyridine in the mixture is obviously seen. Mole fraction of 3-acetylpyridine ( $\times 10^{-5}$ )

Temperature (K)	Pressure (Bar)	Mole fraction of 3-acetylpyridine ( $\times 10^{-5}$ )
313.15	10	3.15
323.15	8	3.15
333.15	6	4.2
343.15	2	0.80
100	120	140
160	180	220
240	260	280

Fig. 2. Mole fraction of 3-acetylpyridine at various pressures and linear regression analysis (experimental data – dot symbols; regression model – dashed lines). In this figure, with fairly small errors below 2%. At pressures below the crossover point, the density effect dominates such that the solubilities decrease with temperature rise (referred as retrograde solubility behavior) [7]. On the other hand, beyond the crossover pressure (in this system is 12 MPa), the solubilities of 3-acetylpyridine in supercritical CO<sub>2</sub> increase both with solvent density and volatility effect in which the latter factor is more prominent.

### 3.2. Semi-empirical density-based correlations

Several

**21** mathematical models have been developed to date for correlating the solubility data

of a particular compound in a particular supercritical solvent

**23** at various pressures and temperatures. Some of these correlation models are highly empirical while some of them have a strong fundamental

theory with thermodynamics consideration. In the present work, the solubility data of 3-acetylpyridine in SC-CO<sub>2</sub> were correlated using semi-empirical density-dependent models and cubic equations of state for evaluating phase equilibria of the system.

**24** Chrastil and Del Valle and Aguilera mathematical models were

used to correlate actual solubilities of 3-acetylpyridine in SC-CO<sub>2</sub> at studied pressures and temperatures. The main advantage of using density-based correlations such as Chrastil and Del Valle and Aguilera is

they do not require any thermophysical properties such as critical point,

**13**acentric factor, vapor **pressure, and molar volume of the**

stud-ied compound [8], which often cannot be easily determined using experimental approach. Chrastil (1982) was the first who devel-oped a semi-empirical equation by assuming the formation of solvato complex between solute and supercritical solvent at equi- librium [9,10]:  $A + kB \leftrightarrow (AB)_k$  (1) where

**17**one molecule of solute **A** associates **with k molecules of** supercritical **solvent B to form one molecule of solvato complex (AB)<sub>k</sub>.**

Chrastil model has mathematical form as follows:  $y = k \exp \left( \frac{a}{T} \right) \rho^b$  (2) The adjustable constants a and b are defined as: ( )  $a = \frac{\Delta H}{R}$  (3)  $b = - \ln \left( \frac{M_A + c}{k M_c} \right) + q$  (4) Fig. 3. Correlations of experimental solubilities of 3-acetylpyridine in supercritical CO<sub>2</sub> (dot symbols) and those predicted with Chrastil (A) and Del Valle and Aguilera (B) mathematical models (wire-mesh). where y is the solubility (kg/m<sup>3</sup>),  $\rho$  is the solvent density (kg/m<sup>3</sup>), k is the average association number,

**6**T is the **absolute temperature (K)**, **R is the universal gas constant (8.314 J mol/K)**, **a is a**

func- tion of

**13**enthalpy of solvation ( $\Delta H_{\text{solv}}$ ) and enthalpy of vaporization ( $\Delta H_{\text{vap}}$ ), **b is a function of average association number**

that depends on the molecular mass of solute ( $M_A$ , 121.14 g/mol) and supercrit- ical solvent ( $M_c$ , 44.01 g/mol),  $\Delta H_{\text{R}}$  is total heat of reaction, which equals to the sum of enthalpy of solvation and enthalpy of vaporization, and q is an adjustable constant. The fitting of experimental solubility data was conducted concurrently at all temperatures using nonlinear least-squares regression technique with

**4**SigmaPlot software package (**Version 12.3, Systat Software Inc.**)

by minimizing root-mean-square devi- ation (RMSD):  $\text{RMSD} = \sqrt{\frac{1}{N} \sum_{i=1}^N (y_{\text{exp}} - y_{\text{calc}})^2}$  (5) Here,  $y_{\text{exp}}$  is the actual solubilities,  $y_{\text{calc}}$  is the calculated solu- bilities, and N is the number of experimental data. The nonlinear least-squares regression technique involves an iterative curve fit- ting procedure to obtain the optimal parameter values that satisfy convergence and tolerance limit. A set of initial values of fitted parameters of the model (i.e., a, b, and k) was determined to start the computation. The number of iterations and step size was designed in such a way to reduce unnecessary computation without affecting the accuracy of predictions. Fig. 3 clearly displays that Chrastil model can visually corre- late experimental solubility data very well with root-mean-square deviation of 0.0027. The optimal fitted parameter values are tabu- lated in Table 3. The adjustable parameter k expresses an average equilibrium association number of which the values of this param- eter are often not an integer based on the fact that the formation of solvato complexes not stoichiometrically takes place, with more or less stable solvato complexes [10]. The value of parameter k obtained from Chrastil model fitting is 3.38, which means that one molecule of 3-acetylpyridine

associates with 3.38 molecules of supercritical CO<sub>2</sub> to form one molecule of solvato complex at equilibrium state. Considering the value of adjustable constant  $a$ , total reaction heat of the dissolution process of 3-acetylpyridine in supercritical CO<sub>2</sub> can be estimated to be  $-56.10$  kJ/mol. The negative value indicates that the dissolution process of 3-acetylpyridine in supercritical CO<sub>2</sub> is exothermic in nature with greater magnitude of energy liberated from the associating process than that required to vaporize solute and solvent molecules by breaking intermolecular forces, characterizing a non-ideal fluid system. Chrastil model has been intensively modified by a number of authors over past few years and one of them is Del Valle and Aguilera. Del Valle and Aguilera (1988) density-based model was chosen because this model compensates for the variation of solute's vaporizing heat with temperature [11]. Del Valle and Aguilera

**4 model has a mathematical form as follows:  $y = k'$**

$\exp(b' + a' T + d' T^2)$  (6)  $H = R a' + 2d' T$  (7) The adjustable parameters  $k'$ ,  $b'$ , and  $a'$  have the same physical meaning to those of Chrastil model. The constant term  $d'/T^2$  is an adjustable parameter associated with the variation of solute's vaporization heat with temperature. Fig. 4 shows that Del Valle and Aguilera model can represent experimental solubility data of 3-acetylpyridine in supercritical CO<sub>2</sub> over the measurement range satisfactorily (RMSD of 0.0026). The optimal fitted parameter values of Del Valle and Aguilera model are presented in Table 3. From the values of adjustable parameters Table 3 The fitted

**24 parameters of Chrastil and Del Valle and Aguilera models for**

correlating system supercritical CO<sub>2</sub> (1) + 3-acetylpyridine (2). Model T (K) Parameters  $a$   $b$   $k$   $d$  RMSD  
 Chrastil 313.15 323.15 333.15 343.15 Del Valle and Aguilera 313.15 323.15 333.15 343.15  $-6746.67$   $-4.50$   
 $3.38$   $-0.0027$   $-16409.25$   $10.09$   $3.39$   $1.59 \times 10^6$   $0.0026$  B. Agustin et al. / Fluid Phase Equilibria 354 (2013) 127–132 131 Fig. 4. Correlation results of mole fraction of 3-acetylpyridine at various pressures and temperatures with

**11 Peng–Robinson equation of state incorporated with quadratic (A) and**

Stryjek–Vera (B) mixing rules.  $a'$  and  $d'$ , one can estimate total reaction heat that ranges between  $-51.76$  kJ/mol and  $-59.16$  kJ/mol. As aforementioned, Del Valle and Aguilera density-dependent model was developed based on the predecessor model (Chrastil equation) therefore this model should give a similar value of fitted parameters. The average association number obtained from Del Valle and Aguilera model fitting is comparable to that of Chrastil model, which means that both models are able to adequately describe the formation of solvato complex between 3-acetylpyridine and supercritical CO<sub>2</sub> at equilibrium. By examining constant  $b'$ , it was found that the calculated values of constant  $q$  for Chrastil and Del Valle and Aguilera equations are considerably different (2.69 vs. 17.32). 3.3. Phase equilibria evaluation using cubic equations of state A number of cubic equations of state with either an empirical, semi-empirical, or theoretical basis are available and have been widely used for phase equilibria evaluation of high-pressure systems and the most popularly used is

**9 Peng–Robinson equation of state. With this approach, the**

vapour–liquid coexistence or supercritical phenomena of a particular fluid mixture can be represented with single equation of state without any conceptual difficulties. However, it should be noted that the applicability of different equations of state is limited and generally accurate over certain measurement range for a

particular system. Peng–Robinson equation of state has a mathematical form:  $P = \frac{RT}{V - b} - \frac{a(T)}{V^2 + 2Vb - b^2}$  (8) where  $P$  is the absolute pressure (atm),  $T$  is the absolute temperature (K),  $V$  is the molar volume (L/mol), and

**6R is the gas constant (8.314 J mol/K).** The adjustable parameters **a** and **b** are both temperature-independent. **For**

single compound system, the constants  $a$  and  $b$  can be determined using a series of equations as follows:  $a(T) = a_c(T_r, \omega) = 0.4572 \frac{R^2 P_c T_c}{T} \alpha(T_r, \omega)$  (9)  $b = 0.0778 \frac{R T_c}{P_c} \omega$  (10) (11) (12) (13) where  $P_c$ ,  $T_c$ ,  $\omega$ , and are critical pressure (MPa), critical temperature (K), acentric factor, and adjustable parameter characteristic for each compound, respectively. The critical properties ( $T_c$  and  $P_c$ ) and acentric factor ( $\omega$ ) of 3-acetylpyridine were predicted using the Joback and Lee–Kesler group contribution methods, respectively with PE2000 computational software (Version 2.9.9a). The prediction of critical pressure and critical temperature using Joback method is simply performed by inputting the number of each functional group and normal boiling point of the compound. This group contribution method has found reasonable accuracy in predicting the critical properties of large variety of compounds [12,13]. Once the critical pressure and critical temperature are obtained, the acentric factor of the compound can be easily determined. To correlate vapour–liquid equilibria of multicomponent systems, appropriate mixing and combining rules should be directly integrated into equations of state parameters. In this regards, both

**15parameters a and b of Peng–Robinson equation of state can be expressed in the**

form of classical van der Waals mixing rule, as shown below [14–16]:  $N a = \sum_{i=1}^N \sum_{j=1}^N x_i x_j a_{ij}$  (14)  $N b = \sum_{i=1}^N \sum_{j=1}^N x_i x_j b_{ij}$  (15)  $\sum_{i=1}^N \sum_{j=1}^N$  In this work, two frequently used mixing rules namely quadratic and Stryjek–Vera (SV) mixing rules were employed for correlating phase equilibrium behavior of 3-acetylpyridine in supercritical CO<sub>2</sub> over the measurement range. The adjustable parameters  $a_{ij}$  and  $b_{ij}$  for quadratic mixing rule can be written in the following mathematical forms [3,9,16]:  $a_{ij} = a_{ii} a_{jj} (1 - k_{ij})$  with  $k_{ji} = k_{ij}$  (16)  $b_{ij} = \sqrt{b_{ii} + b_{jj}} (1 - l_{ij})$  with  $l_{ji} = l_{ij}$  (17) while for Stryjek–Vera mixing rule, the constants  $a$  and  $b$  are given as follows [17]:  $a_{ij} = a_{ii} a_{jj} (1 - x_{ik} k_{ji} - x_{jk} k_{ij})$  with  $k_{ji} \neq k_{ij}$  (18)  $b_{ij} = \sqrt{b_{ii} + b_{jj}} (1 - l_{ij})$  (19)  $k_{ji} = k_{ij} - l_{ij}$  (20) where  $k_{ij}$ ,  $k_{ji}$ , and  $l_{ij}$  are all binary interaction parameters and can be determined by iterative curve fitting procedure of experimental data. For this purpose, PE2000 software was applied to obtain Table 4 Optimal values of

**11binary interaction parameters of Peng–Robinson equation of state**

incorporated with quadratic and Stryjek–Vera mixing rules. Mixing rules T (K) Binary interaction parameters  
 RMSD  $k_{ij}$   $l_{ij}$   $k_{ji}$   $l_{ji}$  Quadratic 313.15 323.15 333.15 343.15 0.544 0.621 0.739 0.932 -0.013 0.016 0.037 0.115  
 - - - - - 0.084 0.071 0.060 0.051 Stryjek–Vera 313.15 323.15 333.15 343.15 0.816 0.521 0.786 0.568  
 0.090 -0.025 0.057 -0.029 0.578 0.608 1.479 0.901 0.238 -0.087 -0.693 -0.333 0.084 0.075 0.062 0.053  
 the optimal values of fitted binary interaction parameters by minimizing objective function and satisfying convergence. The optimal values of

**11binary interaction parameters of Peng–Robinson equation of state**

incorporated with quadratic and Stryjek–Vera mixing rules are tabulated in Table 4. Fig. 4 displays the correlation results between predicted and experimental solubility data of 3-acetylpyridine in supercritical CO<sub>2</sub> using Peng–Robinson equation of state with quadratic and Stryjek–Vera

15 **mixing rules. The root-mean-square deviation between actual and predicted values was in the**

range of 0.051 to 0.084 for quadratic mixing rule, which is comparable with Stryjek–Vera mixing rule. The interaction parameter values of  $k_{ij}$  and  $l_{ij}$  for quadratic mixing rule (Table 4) increase at higher temperature. This result indicates that higher temperature facilitates interaction between “like” molecules, for example solute–supercritical solvent interaction to form a solvato complex. On other hand, the values of binary parameters  $k_{ij}$ ,  $l_{ij}$ , and  $ij$  for Stryjek–Vera mixing rule show inconsistency with temperature variation. The value of  $k_{ij}$  at temperature of 333.15 K is larger than unity, which gives the values of  $a$  and  $a_{ij}$  less than zero. Regardless of similarity of RMSD values, it can be deduced that

9 **Peng–Robinson equation of state incorporated with quadratic mixing rule**

is superior and more sensitive in describing the dependency of binary interaction parameters with temperature than PR-EoS incorporated with Stryjek–Vera mixing rule. 4. Conclusions New

2 **solubility data of 3-acetylpyridine in supercritical carbon dioxide**

over wide range of pressures and temperatures have experimentally measured in this work under static mode. The crossover phenomenon was observed at pressure of 12 MPa and the solubilities of solute in supercritical solvent increased with increasing both pressure and temperature of the system. For mathematical modeling, semi-empirical density-dependent equations namely Chrastil and Del Valle and Aguilera equations gave satisfactory agreement for solubilities correlation over the measurement range. The exothermicity nature of associating process and the equilibrium formation of solvato complexes between 3-acetylpyridine and supercritical carbon dioxide could be consistently described by both equations. The evaluation of phase equilibria of supercritical system using

9 **Peng–Robinson equation of state incorporated with quadratic mixing rule**

show that temperature was the influential factor toward adjustable binary interaction parameters while PR-EoS incorporated with Stryjek–Vera mixing rule failed to describe this behavior. Acknowledgment The first author (Bely Agustin)

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8 **Appendix A. Supplementary data Supplementary material related to this article can be found, in the online version, at <http://dx.doi.org/10.1016/j.fluid>**

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