



PCCP by Suryadi Ismadji

From Hippo (Hippo-03)

Processed on 21-Aug-2018 08:36 WIB

ID: 991684344

Word Count: 6826

Similarity Index	Similarity by Source
9%	Internet Sources: 5% Publications: 8% Student Papers: 1%

sources:

1

1% match (publications)

[Mónica Hernández Rodríguez, Jan Yperman, Robert Carleer, Jens Maggen et al. "Adsorption of Ni\(II\) on spent coffee and coffee husk based activated carbon", Journal of Environmental Chemical Engineering, 2018](#)

2

1% match (publications)

[Iovino, Pasquale, Alessandro Erto, Sante Capasso, Michele Di Natale, Silvana Canzano, Annalisa Lama, and Dino Musmarra. "Experimental analysis of benzene derivative adsorption in single and binary systems using activated carbon", International Journal of Environment and Waste Management, 2015.](#)

3

1% match (publications)

[Sellaoui, Lotfi, Hanen Guedidi, SarraWjhi SarraWjhi, Laurence Reinert, Salah Knani, Laurent Duclaux, and Abdelmottaleb Ben Lamine. "Experimental and theoretical studies of adsorption of ibuprofen on raw and two chemically modified activated carbons: new physicochemical interpretations", RSC Advances, 2016.](#)

4

1% match (publications)

[Mohamed Bouzid, Lotfi Sellaoui, Mohamed Khalfaoui, Hafedh Belmabrouk, Abdelmottaleb Ben Lamine. "Adsorption of ethanol onto activated carbon: Modeling and consequent interpretations based on statistical physics treatment", Physica A: Statistical Mechanics and its Applications, 2016](#)

5

1% match (Internet from 13-May-2018)

<https://www.library.caltech.edu/eds?query=%22Edi+Soetaredjo%2C+Felycia%22&search=y&searchfield=AU&type=AR>

6

1% match (publications)

[Marwa Ben Manaa, Nadia Bouaziz, Bruno Schmalz, François Tran Van, Abdelmottaleb Ben Lamine. "Study of the effect of variation in temperature and pH on the adsorption process of natural gardenia yellow dye into TiO 2 mesoporous for dye sensitized solar cells using the statistical physics formalism: Physicochemical and thermodynamic investigation", Microporous and Mesoporous Materials, 2018](#)

7

1% match (publications)

[Yang, Shitong, Xuemei Ren, Guixia Zhao, Weiqun Shi, Gilles Montavon, Bernd Grambow, and Xiangke Wang. "Competitive sorption and selective sequence of Cu\(II\) and Ni\(II\) on montmorillonite: Batch, modeling, EPR and XAS studies", Geochimica et Cosmochimica Acta, 2015.](#)

8

< 1% match (Internet from 24-Jan-2018)

http://henkelmanlab.org/pubs/evans17_30578.pdf

9

< 1% match (publications)

[Glaysdon Simões dos Reis, Eder Cláudio Lima, Carlos Hoffmann Sampaio, Fabiano Severo Rodembusch et al. "Novel kaolin/polysiloxane based organic-inorganic hybrid materials: Sol-gel synthesis, characterization and photocatalytic properties", Journal of Solid State Chemistry, 2018](#)

10

< 1% match (publications)

[Meili Yin, Zhenhua Li, Zhen Liu, Xinjian Yang, Jinsong Ren. "Magnetic Self-Assembled Zeolite Clusters for Sensitive Detection and Rapid Removal of Mercury\(II\)", ACS Applied Materials & Interfaces, 2012](#)

11

< 1% match (Internet from 04-Jan-2018)

http://www.ijcce.ac.ir/reviewer?_action=info

12

< 1% match (publications)

[Soenjaya, Soegiarto Adi, Nova Handoyo, Felycia Edi Soetaredjo, Artik Elisa Angkawijaya, Yi-Hsu Ju, and Suryadi Ismadji. "Preparation of carbon fiber from water hyacinth liquid tar", International Journal of Industrial Chemistry, 2014.](#)

13

< 1% match (publications)

[Jing Zhang, Nan Chen, Zheng Tang, Yang Yu, Qili Hu, Chuanping Feng. "A study of the mechanism of fluoride adsorption from aqueous solutions onto Fe-impregnated chitosan", Physical Chemistry Chemical Physics, 2015](#)

14

< 1% match (publications)

[Kurniawan, A.. "Performance of durian shell waste as high capacity biosorbent for Cr\(VI\) removal from synthetic wastewater", Ecological Engineering, 201106](#)

15

< 1% match (publications)

[Almasian, Arash, Mazeyar Parvinzadeh Gashti, Mohammad Ebrahim Olya, and Ghazaleh Chizari Fard. "Poly\(acrylic acid\)-zeolite nanocomposites for dye removal from single and binary systems", Desalination and Water Treatment, 2015.](#)

16

< 1% match (publications)

[Sitko, Rafal, Edyta Turek, Beata Zawisza, Ewa Malicka, Ewa Talik, Jan Heimann, Anna Gagor, Barbara Feist, and Roman Wrzalik. "Adsorption of divalent metal ions from aqueous solutions using graphene oxide", Dalton Transactions, 2013.](#)

17

< 1% match (Internet from 25-Jul-2018)

<https://reg.livebyglevents.com/Download.axd?FileID=9b986587-52ff-4a43-be4f-26a05cb329c4>

18

< 1% match (Internet from 05-Aug-2017)

<https://www.cheric.org/PDF/KJChE/KC32/KC32-5-0896.pdf>

19

< 1% match (student papers from 17-Jun-2013)

[Submitted to University of Florida on 2013-06-17](#)**paper text:**

PCCP PAPER

8 **Cite this: Phys. Chem. Chem. Phys., 2017, 19, 25927 Received 5th June 2017, Accepted 6th September 2017 DOI: 10.1039/c7cp03770h rsc.li/pccp**

5 **New insights into single-compound and binary adsorption of copper and lead ions on a treated sea mango shell: experimental and theoretical studies**

Lotfi **Sellaoui**, *a Felycia **Edi Soetaredjo**, b Suryadi **Ismadji**,

b Éder Cláudio Lima,^c Guilherme L. Dotto,^d Abdelmottaleb Ben Laminea and Alessandro Ertoe Herein, adsorption isotherms of Pb(II) and Cu(II) ions on treated sea mango fruit in both single-compound and binary systems were experimentally realized at different temperatures in the range of 30–50 °C. Experimental results show that adsorption of Pb(II) was more as compared to that of Cu(II) ions; however, for both ions, a significant reduction in the adsorption capacity was observed

15 **in the binary system as compared to** that in the **single-**

compound systems. Moreover, under all the investigated conditions, adsorption seems to be promoted by an increase in temperature. To understand and interpret the experimental evidences, the Hill and competitive Hill models developed on the basis of the grand canonical ensemble were applied for the analysis of adsorption equilibrium data. These models contain some physicochemical parameters that allow an exhaustive analysis of the dynamics of single-compound and binary adsorptions. Based on the fitting results, in particular, through the evaluation of the number of ions bonded per site (n and n_i), it was found that lead and copper ions interacted by inclined and horizontal positions on treated sea mango in

2 **single-compound and binary systems**, respectively. In addition, **based on** the same parameters, **a**

significant interaction between ions was retrieved. A study focused on the saturation adsorption capacity in single-compound and binary systems affirmed that the adsorbent was more selective for lead than for copper. The reduction of the adsorbed capacity ratio between the binary and single-compound systems (i.e. Q_b/Q_s) explained and confirmed that an inhibition effect between copper and lead ions at the same receptor

site occurred. Finally, based on the energetic investigations, it was deduced that the adsorption energy represented the dominant factor promoting the greater adsorption of lead than that of copper in both systems. 1. Introduction Water pollution is one of the biggest problems that humans face today. In some areas of the world, the clean water supplies for human consumption have become scarcer and scarcer due to the accidental presence of different pollutants discharged by

3a Unité de Recherche de Physique Quantique, UR 11 ES 54, Université de Monastir, Faculté des Sciences de Monastir, 5000 Monastir, Tunisia. E-mail:

sellaouilofti@yahoo.fr b Department of Chemical Engineering, Widya Mandala Surabaya Catholic University, 601-14 Surabaya, Indonesia c Institute of Chemistry, Federal University of Rio Grande do Sul, UFRGS,

9Porto Alegre, RS, Brazil d Environmental Processes Laboratory, Chemical Engineering Department, Federal University of Santa Maria–UFSM, 1000, Roraima Avenue, 97105-900 Santa Maria, RS, Brazil

e

11Dipartimento di Ingegneria Chimica, dei Materiali e della Produzione Industriale, Università di Napoli Federico II, P. le Tecchio, 80, 80125 Napoli, Italy

industrial or human activities. Industrial activities produce and release a large amount of wastewater during the processing of raw materials into valuable products. Most of the industrial wastewater contains toxic chemicals; therefore, a proper waste- water treatment facility is required to remove hazardous sub- stances from the wastewater before discharging it. 1 Heavy metals are among the most toxic chemicals usually found in wastewater. In general, they are highly threatening to living organisms due to their bio-accumulative behaviour, high toxicity, persistence in nature, and non-biodegradability. The available conventional technologies for the removal of heavy metals from industrial effluents are chemical precipitation, coagulation, reverse osmosis, electro dialysis, ultra-filtration, bio-remediation, phytoremediation, solvent extraction, and adsorption. 2–6 Among the available methods, adsorption is a cost-effective, simple, reliable, and environmentally benign process for the elimination of heavy metals from water and wastewater. While dealing with the adsorption processes, the selection of proper adsorbents is very crucial since the adsorbent determines the effectiveness and inexpensiveness of an adsorption system and overall operation. 7,8 In recent years, the search for renewable and low-cost adsorbents has become a hot topic in adsorption studies. Various ligno- cellulosic materials have been employed as alternative adsor- bents for

17heavy metal removal from aqueous solutions. 9 In this study, we utilized the

sea mango (*Cerbera manghas*) fruit

19 **as a new adsorbent for the removal of Cu(II) and Pb(II) ions from an aqueous**

solution. Sea mango is a small evergreen coastal tree that can grow more than 10 metres in height. Sea mango trees can be easily found in most parts of Indonesia. Since the chemical components of dry sea mango fruits are lignin, cellulose, and hemicellulose, after pre-treatment, this ligno-cellulosic material can be utilized as an alternative adsorbent for heavy metal removal from water or wastewater. Both the single-compound and binary adsorption studies are relevant for the design, optimization, and operation of real water purification plants. In particular, the adsorption of heavy metal ions in multicomponent systems often appears as a strong antagonistic removal process since the properties and concentrations of ions can significantly modify the performances of adsorbents.^{10–12} Since the operating conditions can be potentially infinite in terms of combination of pollutant concentrations, the availability of reliable models to describe and/or predict the behavior of a certain adsorption system is an invaluable tool. In the literature, different models, such as the Langmuir model,^{13,14} modified competitive Langmuir model,¹³ extended Langmuir model,^{15,16} and Freundlich and extended Freundlich model,¹⁵ have been reported to theoretically describe the adsorption phenomena

2 **in both single-compound and binary systems.** Except for **the**

Langmuir model, all these models are empirical and take into account some descriptive aspects of adsorption, thus providing only a phenomenological explanation of the dynamics of the process. In particular, some fundamental details that define the way the adsorbates interact with the adsorbent surface and also with each other in multicomponent systems are rarely considered. For instance, the knowledge of the number of ions bonded per active site and the exact adsorbate position on the adsorbent surface with respect to the molecule shape can provide a direct explanation of the experimental adsorption capacity and the occurrence of any competitive effect, with full physical meaning. In the present study, Hill and competitive Hill models, developed using an analytical theory mainly based on statistical physics,¹⁷ were applied to experimental adsorption isotherms of lead and copper on sea mango shells at different temperatures to provide new and interesting interpretations at the molecular level. The two statistical physics models allow the estimation of the number of ions bonded per active site, the adsorbate position, the density of the receptor sites, the saturation adsorbed quantity, and the molar adsorption energy. All the parameters were interpreted as functions of temperature and provided a deeper description of the single-compound and binary adsorption system. A new parameter based on the investigation of the number of ions bonded per site was also described to explain the inhibition effect arising in binary systems. 2. Experimental Herein, a new adsorbent was produced from sea mango (*Cerbera manghas*) fruits. Raw fruits were obtained from the area near Widya Mandala Surabaya Catholic University, East Java, Indonesia. The fruits were dried in a force circulation oven (Memmert) until the moisture content was around 10%. Subsequently, the dried sea mango fruits were pulverized until their particle size was around 80/100 mesh. Delignification of the sea mango powder was conducted using a sodium hydroxide solution (2 M) at 60 °C for 3 h under constant stirring. After the delignification process was completed, the solid product was separated from the solution, washed several times with water treated by reverse osmosis, and dried at 105 °C for 24 h. The pore characteristics of the pristine sea mango powder and delignified sea mango powder were measured using the nitrogen adsorption method. The measurements were conducted at the boiling point of nitrogen gas (?)

196 1C) using Micromeritics ASAP 2010. Prior to the adsorption/ desorption measurements, the samples were degassed under vacuum conditions at 150 1C. The BET surface area and pore volume of the pristine sea mango powder were 3.28

10 m² g⁻¹ and 0.01 cm³ g⁻¹, respectively. After the delignification process, the surface area

and pore volume of the sea mango powder increased to 15.74

10 m² g⁻¹ and 0.04 cm³ g⁻¹, respectively. The BET surface area

of sea mango is higher as compared to that of other biomasses such as rice straw (1.23 m² g⁻¹), fermented rice straw (6.31 m² g⁻¹),¹⁸ Cucumis melo (8.40 m² g⁻¹),¹⁹ and kraft lignin (0.75 m² g⁻¹).²⁰ The pH_{pzc} value of the sea mango was measured by a zeta meter (Zeta Potential Analyzer, Brookhaven 90Plus). The FTIR analysis of the pristine sea mango and metal-loaded sea mango was conducted using FTIR Shimadzu 8400S. The KBr method was employed to obtain the FTIR spectra. The FTIR spectra were obtained in the range from 4000 to 500 cm⁻¹ (mid-IR wavenumber). The XRD analysis was conducted using a Philips X'pert X-ray diffractometer, and the patterns

12 were acquired at 40 kV and 30 mA

using monochromatic high-intensity CuK α 1 (λ = 0.15405 nm). The SEM images and EDX (energy dispersive X-ray) spectra of the sea mango powder and metal-loaded sea mango were obtained using JEOL JSM-6500F. Before the

12 analysis, the solid samples were coated with a thin layer of platinum under an

argon atmosphere. For single-compound adsorption, solutions of heavy metals at an initial concentration of 200 mg L⁻¹ (for both Cu

18 (II) and Pb (II)) were prepared by dissolving a known amount of

either CuSO₄ · 5H₂O or Pb(C₂H₃O₂)₂ · 3H₂O in reverse osmosis (RO) water. The pH value of the solution was adjusted to 5. Single-compound adsorption experiments were carried out at three temperatures (30, 40, and 50 1C). To describe a complete adsorption isotherm, different masses of adsorbent (sea mango powder) were added to

14 a series of Erlenmeyer flasks containing 100 mL of heavy metal solution. The Erlenmeyer flasks were then placed in a shaker water bath (Memmert) and shaken

until equilibrium conditions were reached (8–10 h). The adsorption isotherms were obtained under equilibrium conditions, and time had no effect on the amount of metal uptake. Subsequently, the adsorbent was removed from the mixture by centrifugation (4000 rpm for 5 min). The residual heavy metal concentration in the solution was analyzed using an atomic absorption spectro- photometer (Shimadzu). The amount of heavy metal adsorbed by the sea mango powder

1 was calculated using the following equation:

$Q_e = \frac{m}{V} \left(\frac{C_0 - C_e}{C_0} \right)$

1 m V (1) where C_0 and C_e are the initial and equilibrium concentrations of heavy metal in the liquid phase (mg L^{-1}), respectively, V is the volume of the heavy metal compound solution (L), and m is the mass of the adsorbent (g). For the

binary system, the heavy metal mixture was prepared by adding known amounts of both $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ and $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$ to 1 L of RO water to reach an initial concentration of both metals equal to 200 mg L^{-1} . We used the same concentration of metals to obtain a better understanding about the competition for the adsorption active sites. Adsorption isotherm experiments were conducted based on the same abovementioned procedure. 3. Characterization of the adsorbent and metal-loaded adsorbent The pH_{pzc} (pH at point zero charge) value of the sea mango was 4.2. At a pH below pH_{pzc} , the surface of the sea mango was positively charged, whereas at a pH above the pH_{pzc} , the surface of the sea mango was negatively charged. The negative surface charge of the sea mango enhances the uptake of metal cations from the solution. For this reason, the

16 adsorption experiments were conducted at pH 5. The FTIR spectra of

the pristine sea mango and Cu and Pb-loaded sea mango are shown in Fig. 1. The adsorption characteristics of heavy metal on cellulosic material were observed at the wavenumbers 3405, 2918, 1375, and 889 cm^{-1} . The broad stretching absorption peak at around 3405 cm^{-1} represents bound $-\text{OH}$ groups. After the adsorption process, this peak shifted to 3423 cm^{-1} for Cu(II) and 3435 cm^{-1} for Pb(II). The band at 2918 cm^{-1} is characteristic of the stretching vibration of $-\text{CH}_2$. For the Cu-loaded sea mango, this band shifted to 2912 cm^{-1} , whereas for the Pb-loaded sea mango, the shift to 2931 cm^{-1} was observed. After exposure to heavy metals, the peaks at 1375 cm^{-1} (CH) and 889 cm^{-1} (non-symmetric out-phase ring) also shifted to 1369 and 901 cm^{-1} for Cu-loaded sea mango and 1363 and 902 cm^{-1} for Pb-loaded sea mango, respectively. The shifting of these wavenumbers indicates that several functional groups are involved in the

adsorption process. A complex formation between some surface functional groups and the metal occurred during the adsorption process. The SEM images of the sea mango and its metal-loaded forms are depicted

16 **in Fig. 2. In general, the adsorption of Cu (II) and Pb (II)**

onto sea mango did not change the surface topography of the adsorbent. The EDX spectra of sea mango and its Fig. 1 FTIR spectra of sea mango and metal-loaded sea mango. metal-loaded forms are also shown in Fig. 3. The EDX spectra of the Cu and Pb-loaded sea mango indicate that a substantial amount of Cu and Pb has been

17 **adsorbed on the surface of the adsorbent. The**

Cu and Pb peaks have been detected after the adsorption process, which indicate that copper and lead ions may attach to some surface functional groups of the sea mango. High percentages of C, O, and N in sea mango and its metal-loaded forms suggest that these elements are the building blocks of biomass. X-ray diffraction patterns of the sea mango before and after Cu(II) and Pb(II) adsorption are shown in Fig. 4. New diffraction peaks were observed for both Cu and Pb-loaded sea mango. For Cu-loaded sea mango, new peaks appear at the 2θ values around 14.111, 27.121, and 30.381, whereas for Pb-loaded sea mango, the peaks appear at the 2θ values around 24.651, 26.601, and 35.311. The appearances of new peaks in the XRD spectra indicate the formation of copper or lead-containing substances. 4. Adsorption isotherms in the single-compound and binary system Methodologically, before interpreting the single and binary adsorption process through the parameters appearing in the Fig. 2 SEM images of (a) sea mango, (b) Cu-loaded sea mango, and (c) Pb-loaded sea mango. statistical physics models, it is important to describe the profile of the adsorption isotherms and explain the corresponding phenomenon. Adsorption isotherms in both the single-compound and binary systems are shown in Fig. 5. Experimentally and based on Fig. 5, the adsorption isotherms show a classical behavior in relation to the monotonic increase of the adsorbed capacity as a function of equilibrium concentration. The increase in this quantity indicates that the treated sea mango adsorbent accepts the necessary number of ions bonded per site, Fig. 3 EDX spectra of sea mango and its metal-loaded forms. and finally, a number of receptor sites are occupied, which in some cases appear to be the maximum possible (asymptote). Moreover, the adsorption data show that the adsorption of Pb(II) is more than that of

7 **Cu(II) in both the single- compound and binary systems.**

This ranking is retained for all the investigated temperatures. It is interesting to observe that in the binary system, the adsorption capacity of both ions significantly decreases to values approximately equal to half of their corresponding values retrieved in the single-compound systems. Moreover, for this evidence, the behavior is the same at all Fig. 4 XRD spectra of sea mango and heavy metal-loaded sea mango. the investigated temperatures; this testifies the occurrence of strong competition phenomena during simultaneous adsorption. Finally, it is evident that in all the investigated systems and under the experimental conditions tested, adsorption is enhanced with the increasing temperature. To theoretically explain the physical phenomena involved in the adsorption of Pb

7(II) and Cu (II) in the single- compound and binary systems,

two statistical physics models¹⁷ were applied, providing new interpretations at the molecular level. The main parameters affecting the single-compound and binary adsorption retrievable from the statistical physics treatment of the experimental data, i.e., the number of ions bonded per site and the adsorption energy, were used for the interpretation of the experimental evidence. 4.1 Models for single-compound and binary adsorption For a complete theoretical description of the single-compound and binary adsorption phenomena, many conventional theories have been developed. Among them, the most tested theories are the Langmuir and competitive Langmuir models. These models were constructed based on easy assumptions, e.g., the receptor sites are energetically identical regardless of the adsorbate considered. Furthermore, they speculate that a localized adsorbed receptor site accommodates one adsorbate per active site. In terms of interactions, Langmuir and competitive Langmuir models neglect any interactions among adsorbates. Despite these models being successfully used in many studies, they present a serious drawback, which is not related to the mathematical form of the equation, but it concerns the assumptions made and their consequence in the interpretation of the results. These simple assumptions produced models in which the adsorption parameters did not have a physical meaning or just provided average values of the properties that did not help to fully understand the adsorption phenomena. To overcome these Fig. 5 Adsorption isotherms

5of copper and lead ions on sea mango shell in single-compound and

binary systems (pH 5, T = 30, 40, and 50 °C). limitations, we have recently developed interesting models for single-compound and binary adsorption

4on the basis of the grand canonical ensemble in statistical physics.

Among these models, Hill¹⁷ and competitive Hill models¹⁷ might have a good impact on this study. The strong difference between the empirical models and the last cited models is that all the parameters of the Hill and competitive Hill models have a defined physical meaning; this is the main advantage of the statistical physics models as compared to that of the empirical models. Additionally, the Hill and competitive Hill models represent the

3general case of the Langmuir and competitive Langmuir models. Indeed, the Langmuir

and competitive Langmuir models assume that each active site of the adsorbent accepts one ion/molecule, whereas the Hill and competitive Hill models assume that each active site of the adsorbent can accept n or n_i ($i = 1, 2$) ion/molecules. Moreover, these parameters vary with temperature; hence, it is worthy to briefly describe the model development. 4.2 Hill model for single adsorption To construct the Hill model, we speculated that the receptor sites of the treated sea mango shell are energetically identical. Based on this energetic consideration, it has been supposed that a receptor site (S) is characterized by a defined (ϵ)

energy,²¹ and it accepts a variable number of ions. When the adsorption saturation condition is reached, a certain number of receptor sites are occupied by lead (A^{2+}) or copper ions (B^{2+}). For this reason, we consider that this number of receptor sites is NM . Chemically, we can describe the single adsorption process by the following pseudo-reactions:²¹ $nA^{2+} + pS \rightleftharpoons A_n^{2+}pS$ $nB^{2+} + pS \rightleftharpoons B_n^{2+}pS$ (2) where $A_n^{2+}pS$ and $B_n^{2+}pS$ represent the formed complex (adsorbate-adsorbent),

n is the number of ions bonded per receptor site

on the adsorbent surface. Based on the grand canonical ensemble definition in statistical physics, the partition function of one identical site is written as follows:^{21,22} $z_{gc} = 1 + e^{b(e+m)}$ (3) Herein, two meaningful terms are included in the partition function: e represents the

adsorption energy of the receptor site, m represents the chemical potential of the receptor site, and b is the Boltzmann factor. As abovementioned, the

partition function is simple and contains only two energetic parameters (e and m); hence, it is reasonable to attain a simple formulation of the Hill model, which expresses the

evolution of the adsorption capacity as a function of heavy metal equilibrium concentration

(c_e), as given by the following expression:^{21,22} $Q = \frac{c_1}{2} \frac{1}{1 + e^{-b(e_1 + m_1)} + e^{-b(e_2 + m_2)}}$ (4) $1/p$ c_e In this expression, Q represents the adsorption capacity of the heavy metal and $c_1/2$ and c_e are the concentrations at half saturation and at equilibrium of heavy metal, respectively. **4.3 Competitive Hill model for binary adsorption** To describe the competitive effect between ions, it is speculated that these ions interact at the same type of receptor site and this site exclusively accepts lead and copper ions.¹⁷ Herein, two energies were considered for the competitive Hill model notation (e_1) and (e_2), describing the interactions between the receptor site and either lead or copper ions, respectively.¹⁷ Moreover, two binary pseudo adsorption reactions explaining the competitive effect between the studied ions can be written as follows:¹⁷ $n_1C^{2+} + pS \rightleftharpoons C_{n_1}^{2+}pS$ $n_2D^{2+} + pS \rightleftharpoons D_{n_2}^{2+}pS$ (5) where C^{2+} and D^{2+} represent Pb^{2+} and Cu^{2+} ions and n_1 and n_2 are the number of ions bonded per receptor site, respectively. The

partition function of the described receptor site can be expressed as

follows:¹⁷ $z_{gc} = 1 + e^{b(e_1 + m_1)} + e^{b(e_2 + m_2)} = T_1 + T_2 + T_3$ (6) Unlike the expression of the single-compound Hill model, this partition function of the competitive Hill model is more complicated because it contains three terms. The first term (T_1) indicates that the receptor site cannot be occupied (empty) and $T_2 + T_3$ explain the occupation of the receptor site by two ions. m_1 and m_2 represent the chemical potentials of Pb^{2+} and Cd^{2+} ions at the receptor site, respectively. Under these hypotheses, the competitive Hill model expressing the evolution of adsorption capacity as a function of heavy metal equilibrium concentrations (c_e)

and c_{e2}) is written as follows: $17 n_1 N M c_{e1} n_1 Q_{a1} \frac{1}{4} c_{01} 1 p c_{e1} n_1 p c_{e2} c_{01} c_{02} Q_{a2} \frac{1}{4} n_2 N M c_{e2} c_{01} c_{02} n_2 n_2 (7) (8)$ where NM is the density of the receptor site; c_{01} and c_{02}

3represent the concentrations **at half saturation of** each adsorbate; c_{e1} **and** c_{e2} are **the**

heavy metal equilibrium concentrations. A simple fitting method

3based on the **Levenberg–Marquardt** algorithm²³ shows a

good correlation between the statistical physics models and the experimental data in

2single-compound and binary systems. Indeed, **the** values of

the correlation coefficient R^2 measuring the fitting quality vary from 0.975 to 0.994 and from 0.971 to 0.992 in the single-compound and binary system, respectively. Additionally, the fitting method provides all the values of the model parameters at different Table 1 Adsorption parameters in

2single-compound and binary systems deduced by **the Hill and**

competitive Hill models Pb(II):single

1330 1C 40 1C 50 1C

Cu(II):single

1330 1C 40 1C 50 1C

Q_{sat} (mg g^{-1}) n NM 1.02 1.44 58.11 64.08 56.97 44.44 1.98 1.01 67.12 39.25 33.83 39.18 1.28 49.44 38.28 1.74 52.17 29.88 Pb(II):binary Cu(II):binary 30 1C 40 1C 50

131C 30 1C 40 1C 50 1C

n_i Q_{sat} (mg g^{-1}) NM 0.50 0.55 26.11 33.12 52.12 60.21 0.38 0.37 42.11 17.12 110.81 46.27 0.28 0.33 22.18 27.21 79.21 81.23 temperatures, as summarized in Table 1. An example of the adsorption isotherm fitting is reported in the Appendix. 5. Hill and competitive Hill model parameter interpretation 5.1

1 Number of ions bonded per site (n and ni) The

n and ni ($i = 1, 2$) parameters balancing the pseudo-reactions (eqn (2) and (5)) are characterized by physical and chemical aspects, which play many important roles in the single- compound and binary adsorption studies. Based on the scientific literature and to the best of our knowledge, the position of copper and lead ions on the adsorbent surface upon adsorption has not described to date regardless of the studied adsorbent. These two parameters are able to provide useful information about the position at each temperature by a systematic method. In our recent study, two interesting cases were described to know

6 the position of the adsorbate (ions/molecules) on the adsorbent surface.

It is a simple method and applicable for any adsorption system. The first case describes a parallel position on the adsorbent if the number of ions bonded per site is lower than 1, i.e., the ions interact with a minimum of two receptor sites (multi-link).^{21–23} The second case assumes a perpendicular position if the n and ni values are higher than 1, i.e., a receptor site is occupied by at least one ion.^{21–23} Based on the fitting results (Table 1), in the single-compound systems, the n parameter values are higher than 1 for all the temperatures and for both adsorbates; however, in the binary adsorption system, the contrary is observed. From these results, it was determined that copper and lead ions were bonded in a perpendicular position on the treated sea mango in single-compound systems, whereas in the binary systems, they were bonded in a parallel position. As an explicative example, the n parameter value of Pb^{2+} at 30 °C in the single-compound system is 1.02, whereas in the binary system, this parameter value is reduced to 0.50. This indicated that Pb^{2+} ion was bonded practically with one receptor site in the single-compound system, but by two receptor sites in the binary system. This result allows for hypothesizing a mechanism of adsorption of two ions on the adsorbent. In fact, an electrostatic interaction between cations and the adsorbent surface appears to be most probable because of the demonstrated competition effects between ions of the same charge and the higher number of receptor sites necessary to bind lead in the binary system. The latter evidence testifies that a preferential adsorption with copper ions occurs. and for this reason, a higher number of receptor sites is necessary to bind one lead ion (i.e., a weaker interaction with lead makes it necessary to bind to more active sites). This conclusion finds support in the energetic characterization of the adsorption systems, as reported hereinafter. The evolution of the n and ni parameters as a function of temperature is depicted in Fig. 6. Based on this evolution, it can be seen that the

1 number of ions bonded per site (in the

single-compound system) has a linear trend with temperature; this reflects that the thermal agitation effect is the dominant factor, which explains this behavior.²⁴ With regard to the evolution of the ni parameter with temperature in the binary system, the behavior is less simple and not monotonic for both ions due to specific physical interactions.

4 When the number of copper ions bonded per site increases, the number of

lead ions bonded per site decreases. This evolution is mainly due to the competitive effect of these ions for the same receptor site. Hence, an inhibition effect is created between these ions, i.e., a partial exclusion. For a possible explanation of this inhibition effect, it is useful to observe that when lead ion (for example) interacts Fig. 6 Temperature effect on the number of ions bonded per site in the single (a) and binary system (b). with two receptors sites, the presence of copper ion can hide these receptor sites. The Hill and competitive Hill models provide explanation of the competition between ions for the same receptor site by introducing a new parameter. This new parameter is defined as follows: n_b/n_s (or n_i/n) where n_s and n_b are the number of ions of the same metal bonded per site

7in the single- compound and binary systems, respectively. In the

systems under analysis, all the values of this ratio were lower than 1. This result indicated that the number of ions bonded per site was reduced from the single to binary adsorption system; this reflected the competitive effect. It was concluded that a visible interaction was established between two ions at the same receptor site. Finally, a comparison with the literature models (e.g. Langmuir and extended Langmuir models) showed interesting differences. In fact, these models assume that the adsorbent active site can bind one ion regardless of the adsorbate nature. In our study, the models suppose that the adsorbent can bind a variable number (n and n_i) of ions in the single-compound and binary systems; this appears to be most representative of the occurring phenomena. Herein, we reiterate that it is very useful to estimate the number of ions bonded per site to facilitate the understating of the adsorption phenomena. 5.2 Density of the receptor sites When saturation is reached, a number of receptor sites are completely occupied (NM). The

3evolution of this parameter with temperature is depicted in Fig.

7 for both metals and both the single-compound and binary systems. It can be clearly observed that the receptor site density is influenced by temperature. For the single adsorption system, when the temperature increases, the density decreases and this is probably due to thermal collisions.²⁴ The opposite behavior was observed in the binary system. The increase of density in the binary system can be explained by the absence of the steric hindrance effect between ions at the same receptor site. When the steric hindrance effect is neglected, it is possible that many new receptor sites appear for the adsorption of the ions. 5.3 Evaluation of adsorption capacity at saturation in the single-compound and binary systems The saturation adsorption capacity represents one of the most meaningful parameters describing the single-compound and binary adsorption systems as it provides invaluable information about the performance limits of an adsorbent. In Fig. 8, the

6Fig. 7 Evolution of density of receptor sites as function

of temperature in single-compound (left) and binary (right) adsorption. Fig. 8 Saturation adsorption capacity of copper and lead ions in the single-compound (left) and binary (right) system as a function of temperature. saturation adsorption capacity is depicted as a function of the single-compound and binary systems were calculated at temperature for all the investigated experimental conditions. different temperatures as follows:¹⁷ A clear dependence of this parameter on temperature is observed in Fig. 6 and is more clearly illustrated in Fig. 8, $e^{-1/4} k_B T$ In cc1=s2 (9) which shows that the adsorption capacity increases when the

temperature increases. This evolution type reflects that the temperature promotes the adsorption of lead and copper ions $ei \frac{1}{4} kBT \ln csi$ (10) on treated sea mango; hence, single-compound and binary c_i adsorption have an endothermic character. It was noted that the adsorption capacity value in the binary system was reduced where csi is the solubility of the heavy metal i , as compared to that of the single-compound counterpart. For \ln Fig. 9, the calculated values for the investigated systems example, for Cu^{2+} , this quantity is 38 mg g^{-1} at $T = 30 \text{ }^\circ\text{C}$ in the as a function of temperature are reported. The adsorption single-compound system, whereas in the binary system, it is energy values varied from 8.66 to $37.14 \text{ kJ mol}^{-1}$ in the reduced to 18 mg g^{-1} at the same temperature. This means single-compound system, whereas in the binary system, they that the Cu^{2+} adsorption capacity was reduced by about 47% . ranged from 3.88 to $16.22 \text{ kJ mol}^{-1}$. This evidence indicates that As abovementioned in Section 5.1, an evident interaction was both the single-compound and binary adsorption processes occur established between two ions at the same receptor site. To describe by physisorption;¹⁷ however, the single-compound adsorption and quantify this interaction effect, it is necessary to determine the systems show high adsorption energy, which is similar to chemi- ratio between the adsorption capacity at saturation in the binary sorption. Based on our reported fitting results, it is clear that the system and the corresponding (i.e. same) ion in the single- adsorbed quantity of Pb^{2+} is greater as compared to that of Cu^{2+} at compound system.²⁵ each temperature and in both the single-compound and binary $Q_b/Q_s = 1$: there is no significant interaction between lead systems. Classically and based on the reported studies, we can and copper ions. explain this difference by many reasons such as the differences in Q_b/Q_s $4 : 1$: adsorption of copper ion is promoted by the Pauling electronegativity²⁶ and the ionic size of the adsorbates.²⁶ presence of lead ion or the inverse. However, apart from the classical interpretations, in light of the $Q_b/Q_s > 1$: the adsorption of copper ion is affected by modelling analysis carried out, it is possible to attribute new copper ions or the inverse. explanations with regard to this behavior. In fact, as previously For both ions, Q_s was higher than Q_b at each temperature; thus, reported, our fitting results in the single-compound systems the ratio was lower than 1 . In conclusion, this result demonstrated showed that the number of ions bonded per Pb^{2+} site was higher that an inhibition effect was established between lead and copper than the number of Cu^{2+} . Despite this number being reduced ions at the same site of the treated sea mango shell. in the binary system, the receptor site of the adsorbent still prevalently accepted Pb^{2+} ions. In parallel, the adsorbed quantity was characterized by a similar behavior in the single-compound and binary system, i.e., the adsorbed quantity of Pb^{2+} was greater 6 . Adsorption energy and comparative than that of Cu^{2+} in single-compound systems, and besides a study reduction, the same behavior was observed in the binary system. Interestingly, the adsorption energy showed the same qualitative The adsorption energy parameter represents a tool to charac- trend with temperature as the adsorbed quantity in the single- terize the interactions of lead and copper ions with the adsor- compound and binary systems. Indeed, the Pb^{2+} adsorption bent surface. Using the two expressions deduced from the energy is greater than the Cu^{2+} adsorption energy, and this energy Hill and competitive Hill models, the adsorption energies for is reduced in the binary system. Reasonably, it can be concluded Fig. 9 Adsorption energy evolution in the single-compound (left) and binary (right) adsorption system as a function of temperature. that this energy is the dominant factor controlling the

2single- compound and binary adsorption systems, favoring the adsorp- tion of

Pb^{2+}

2in both single-compound and binary systems. In conclusion, the reduction of

the adsorbed quantity in the binary system is mainly due to the reduction in the adsorption energy. 7. Conclusion Herein, the

18 **adsorption of Pb (II) and Cu (II) ions on the treated sea mango was**

studied as a function of the temperature

2 **in both single-compound and binary systems.**

2 **Experimental results** showed **that the adsorption of**

Pb(II) was more than that of

7 **Cu(II) ions in both the single- compound and binary systems.**

Simultaneously, both ions experienced a significant reduction in the adsorption capacity

15 **in the binary system as compared to** those in the **single- compound system.**

Finally, under all **the**

investigated conditions, adsorption seems to be promoted by an increase in temperature. To provide a meaningful explanation for this experimental evidence, a dedicated modelling the single or binary system, respectively, based on investigation of the number of ions bonded per site. The study of the same parameters explained that an inhibition effect appeared between ions for the same receptor site in the binary system with respect to the single-compound counterpart. The interpretation of the adsorbed quantity confirmed this inhibition effect and explained that the adsorbent was more selective for lead ions. Among the adsorption parameters, adsorption energy was found to be most significant as it represented the main factor controlling the extent of single-compound and binary adsorption.

8 **Conflicts of interest There are no conflicts to declare.**

Appendix Example of adsorption isotherm fitting in the single and binary systems. analysis was carried out and models, based on statistical References physics, were derived. In particular, the Hill and competitive Hill models devel- 1 P. Iovino, S. Canzano, S. Capasso, A. Erto and D. Musmarra, oped using the statistical physics were applied to provide new Chem. Eng. J., 2015, 227, 360–367. and interesting interpretations of lead and copper ion adsorp- 2 M. Benjamin, Water Chemistry, New York, 2nd edn, 2002. tion on treated sea mango in the single-compound and binary 3 F. Fenglian and Q. Wang, J. Environ. Manage., 2011, 92, systems. Using the model parameters, the adsorbate position of 407–418. each bonded ion was described, and it was shown that the ions 4 S. Satyro, M. Race, Di. F. Natale, A. Erto, M. Guida and interacted by either

the perpendicular or parallel position in R. Marotta, *Chem. Eng. J.*, 2016, 283, 1484–1493. 5 G. J. D. M. Rocha, V. M. Nascimento, A. R. Gonçalves, V. F. N. Silva and C. Martín, *Ind. Crops Prod.*, 2015, 64, 52–58. 6 Y. Pan, P. Cai, M. Farmahini-Farahani, Y. Li, X. Hou and H. Xiao, *Appl. Surf. Sci.*, 2016, 358, 333–340. 7 A. Bhatnagar and M. Sillanpaa, *Chem. Eng. J.*, 2010, 157, 277–296. 8 A. Erto, L. Giraldo, A. Lancia and J. C. M. Piraján, *Water, Air, Soil Pollut.*, 2013, 224(4), 1531–1541. 9 S. Babel and T. A. Kurniawan, *J. Hazard. Mater.*, 2003, 97, 219–243. 10 A. Erto, D. F. Natale, D. Musmarra and A. Lancia, *Adsorption*, 2015, 21(8), 611–621. 11 F. Bouhamed, Z. Elouear, J. Bouzid and B. Ouddane, *Environ. Sci. Pollut. Res.*, 2016, 23, 15801–15806. 12 J. H. Park, O. S. Sik, S. H. Kim, J. S. Cho, R. D. Delaune and D. C. Seo, *Chemosphere*, 2016, 142, 77–83. 13 E. F. Soetaredjo, A. Kurniawan, O. L. Ki and S. Ismadji, *Chem. Eng. J.*, 2013, 219, 137–148. 14 H. Muhamad, H. Doan and A. Lohi, *Chem. Eng. J.*, 2010, 158, 369–377. 15 E. F. Soetaredjo, S. Ismadji, S. P. Santoso, O. L. Ki, A. Kurniawan and Y. H. Ju, *Chem. Eng. J.*, 2013, 231, 406–413. 16 A. S. Luna, A. L. H. Costa, A. C. A. da. Costa and C. A. Henriques, *Bioresour. Technol.*, 2010, 101, 5104–5111. 17 L. Sellaoui, M. Bouzid, L. Duclaux, L. Reinert, S. Knani and A. B. Lamine, *RSC Adv.*, 2016, 6, 67701–67714. 18 J. Y. Wang, H. Cui, C. W. Cui and D. F. Xing, *Ecol. Eng.*, 2017, 95, 793–799. 19 S. T. Akar, S. Arslan, T. Alp, D. Arslan and T. Akar, *Chem. Eng. J.*, 2012, 185–186, 82–90. 20 M. B. Sciban, M. T. Klasnja and M. G. Antov, *Ecol. Eng.*, 2011, 37, 2092–2095. 21 L. Sellaoui, H. Guedidi, S. Knani, L. Reinert, L. Duclaux and A. B. Lamine, *Fluid Phase Equilib.*, 2015, 387, 103–110. 22 L. Sellaoui, N. Mechi, É. C. Lima, G. L. Dotto and A. B. Lamine, *J. Phys. Chem. Solids*, 2017, 109, 117–123. 23 L. Sellaoui, H. Guedidi, S. Masson, L. Reinert, J. M. Levêque, S. Knani, A. B. Lamine, M. Khalfaoui and L. Duclaux, *Fluid Phase Equilib.*, 2016, 414, 156–163. 24 L. Sellaoui, G. L. Dotto, J. O. Gonçalves, L. A. A. Pinto, S. Knani and A. B. Lamine, *J. Mol. Liq.*, 2016, 222, 151–158. 25 F. Qin, B. Wen, X. Q. Shan, Y. N. Xie, T. Liu, S. Z. Zhang and S. U. Khan, *Environ. Pollut.*, 2006, 144, 669–680. 26 T. Depci, A. R. Kul and Y. Önal, *Chem. Eng. J.*, 2012, 200–202, 224–236.

Paper PCCP PCCP Paper This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25927 25928 | *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 This journal is © the Owner Societies 2017 This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25929 25930 | *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 This journal is © the Owner Societies 2017 This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25931 25932 | *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 This journal is © the Owner Societies 2017 This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25933 25934 | *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 This journal is © the Owner Societies 2017 This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25935 25936 | *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 This journal is © the Owner Societies 2017 This journal is © the Owner Societies 2017 *Phys. Chem. Chem. Phys.*, 2017, 19, 25927–25937 | 25937