



Turnitin Originality Report

10.1007_s13762-012-0114-y by Suryadi Ismadji

From paper 2 (Hippo-hippo 02)

Processed on 16-Feb-2018 07:30 WIB

ID: 916705877

Word Count: 6624

Similarity Index	Similarity by Source
20%	Internet Sources: 9% Publications: 17% Student Papers: 6%

sources:

- 1

 2% match (student papers from 14-Jan-2014)
[Submitted to TechKnowledge on 2014-01-14](#)

- 2

 1% match (publications)
[Ullah, Zia, Sajjad Hussain, Saima Gul, Sabir Khan, and F.K. Bangash. "Use of HCl-modified bentonite clay for the adsorption of Acid Blue 129 from aqueous solutions", Desalination and Water Treatment, 2015.](#)

- 3

 1% match (Internet from 21-Oct-2014)
<http://www.soci.org/News/Separation-Science/~media/Files/Conference%20Downloads/2008/IEX%20Jul%2008/Carmonasecure.ashx>

- 4

 1% match (publications)
[Chandra, T.C.. "Adsorption of basic dye onto activated carbon prepared from durian shell: Studies of adsorption equilibrium and kinetics", Chemical Engineering Journal, 20070301](#)

- 5

 1% match (publications)
[S. Ismadji, S. K. Bhatia. "A Modified Pore-Filling Isotherm for Liquid-Phase Adsorption in Activated Carbon", Langmuir, 2001](#)

- 6

 1% match (student papers from 04-Sep-2015)
[Submitted to Universidad Estadual Paulista on 2015-09-04](#)

- 7

 1% match (publications)
[Elroy Nathaniel, Alfin Kurniawan, Felycia Edi Soeteredjo, Suryadi Ismadji. "Organo-bentonite for the adsorption of Pb\(II\) from aqueous solution: Temperature dependent parameters of several adsorption equations", Desalination and Water Treatment, 2012](#)

- 8

 1% match (publications)
["New Penicillins Research from A.K. Rahardjo and Co-Researchers Described.", Biotech Week, August 3 2011 Issue](#)

- 9

 1% match (publications)

[Ismadji, Suryadi, Felycia Edi Soetaredjo, and Aning Ayucitra. "The Kinetic Studies in the Adsorption of Hazardous Substances Using Clay Minerals", SpringerBriefs in Molecular Science, 2015.](#)

10

1% match (publications)

[Soetaredjo, F.E.. "KOH/bentonite catalysts for transesterification of palm oil to biodiesel", Applied Clay Science, 201108](#)

11

1% match (publications)

[M. Erdemoğlu, S. Erdemoğlu, F. Sayılkan, M. Akarsu, Ş. Şener, H. Sayılkan. "Organo-functional modified pyrophyllite: preparation, characterisation and Pb\(II\) ion adsorption property", Applied Clay Science, 2004](#)

12

1% match (publications)

[SHAO-LONG DU. "DEACIDIFICATION OF ADLAY SEED \(COIX LACHRYMA-JOBI VAR. MAYUEN\) MISCELLA WITH ANION-EXCHANGE RESIN", Journal of Food Process Engineering, 12/2007](#)

13

< 1% match (publications)

[Song, Hua, Gang Yang, HuaLin Song, XueHan Cui, Feng Li, and DanDan Yuan. "Kinetic and thermodynamic studies on adsorption of thiophene and benzothiophene onto AgCeY Zeolite", Journal of the Taiwan Institute of Chemical Engineers, 2016.](#)

14

< 1% match (publications)

[Herman Hindarso. "Adsorption of Benzene and Toluene from Aqueous Solution onto Granular Activated Carbon", Journal of Chemical & Engineering Data, 07/2001](#)

15

< 1% match (publications)

[Ismadji, Suryadi, Felycia Edi Soetaredjo, and Aning Ayucitra. "Natural Clay Minerals as Environmental Cleaning Agents", SpringerBriefs in Molecular Science, 2015.](#)

16

< 1% match (publications)

[Southichak, B.. "Phragmites australis: A novel biosorbent for the removal of heavy metals from aqueous solution", Water Research, 200607](#)

17

< 1% match (publications)

[Hartono, Chynthia Devi, Kevin Jonathan Marlie, Jindrayani Nyoo Putro, Felycia Edi Soetardjo, Yi Hsu Ju, Dwi Agustin Nuryani Sirodj, and Suryadi Ismadji. "Levulinic acid from corncob by subcritical water process", International Journal of Industrial Chemistry, 2016.](#)

18

< 1% match (Internet from 07-Mar-2016)

http://is.pcz.pl/static/pdf/2013/zeszyt3/3_3-16_2013_DabekKusmierrek.pdf

19

< 1% match (publications)

[Agustin, Bely, Shi-Yow Lin, Alfin Kurniawan, Yi-Hsu Ju, Felycia Edi Soetaredjo, and Suryadi Ismadji. "Solubilities of 3-acetylpyridine in supercritical carbon dioxide at several temperatures and pressures: Experimental and modeling", Fluid Phase Equilibria, 2013.](#)

20

< 1% match (publications)

[Sharma, Ponchami, Binoy K. Saikia, and Manash R. Das. "Removal of methyl green dye molecule from aqueous system using reduced graphene oxide as an efficient adsorbent: Kinetics, isotherm and thermodynamic parameters", Colloids and Surfaces A Physicochemical and Engineering Aspects, 2014.](#)

21

< 1% match (publications)

[Bilge Erdem. "Preparation of HDTMA-bentonite: Characterization studies and its adsorption behavior toward dibenzofuran", Surface and Interface Analysis, 2010](#)

22

< 1% match (student papers from 09-Nov-2016)

[Submitted to Thapar University, Patiala on 2016-11-09](#)

23

< 1% match (publications)

[Soenjaya, Soegiarto Adi, Nova Handoyo, Felycia Edi Soetaredjo, Artik Elisa Angkawijaya, Yi-Hsu Ju, and Suryadi Ismadji. "Preparation of carbon fiber from water hyacinth liquid tar", International Journal of Industrial Chemistry, 2014.](#)

24

< 1% match (student papers from 05-May-2016)

[Submitted to University of Nottingham on 2016-05-05](#)

25

< 1% match (student papers from 14-Jun-2012)

[Submitted to Universiti Sains Malaysia on 2012-06-14](#)

26

< 1% match (student papers from 08-Sep-2011)

[Submitted to The Robert Gordon University on 2011-09-08](#)

27

< 1% match (Internet from 02-Mar-2016)

https://ncsu.edu/bioresources/BioRes_06/BioRes_06_2_2161_c_Hubbe_Removal_Pollutants_Review.pdf

28

< 1% match (Internet from 06-Nov-2017)

<http://nardus.mpn.gov.rs/bitstream/handle/123456789/8336/Disertacija.pdf?isAllowed=y&sequence=1>

29

< 1% match (Internet from 16-Sep-2017)

https://repositorio.unesp.br/bitstream/handle/11449/97884/bastos_ac_me_ilha.pdf?isAllowed=y&sequence=1

30

< 1% match (publications)

[Kusuma, Ricky Indra, Johan Prabowo Hadinoto, Aning Ayucitra, Felycia Edi Soetaredjo, and Suryadi Ismadji. "Natural zeolite from Pacitan Indonesia, as catalyst support for transesterification of palm oil", Applied Clay Science, 2012.](#)

31

< 1% match (publications)

[Arum Adriani Liman. "Potential application of waste neem leaves for bleaching of low-quality crude palm oil", *Asia-Pacific Journal of Chemical Engineering*, 2010](#)

32

< 1% match (publications)

[Do, . "Fundamentals of Pure Component Adsorption Equilibria", *Series on Chemical Engineering*, 1998.](#)

33

< 1% match (publications)

[Ismadji, Suryadi, Felycia Edi Soetaredjo, and Aning Ayucitra. "Modification of Clay Minerals for Adsorption Purpose", *SpringerBriefs in Molecular Science*, 2015.](#)

34

< 1% match (publications)

[Soetaredjo, Felycia Edi, Alfin Kurniawan, Ong Lu Ki, and Suryadi Ismadji. "Incorporation of selectivity factor in modeling binary component adsorption isotherms for heavy metals-biomass system", *Chemical Engineering Journal*, 2013.](#)

35

< 1% match (Internet from 28-Dec-2017)

<http://www.philrice.gov.ph/wp-content/uploads/2017/12/Biology-and-management-of-invasive-apple-snails.pdf>

36

< 1% match (publications)

[Sulaymon, Abbas H., Saib A. Yousif, and Mustafa M. Al-Faize. "Competitive Biosorption of Lead Mercury Chromium and Arsenic Ions onto Activated Sludge in Batch Adsorber", *Aquatic Science and Technology*, 2012.](#)

37

< 1% match (publications)

[Mohan, D.. "Single, binary and multi-component adsorption of copper and cadmium from aqueous solutions on Kraft lignin-a biosorbent", *Journal of Colloid And Interface Science*, 20060515](#)

38

< 1% match (publications)

[Khalid Z. Elwakeel. "Adsorption of malathion on thermally treated egg shell material", *Water Science & Technology*, 02/2010](#)

39

< 1% match (publications)

[Kurniawan, A.. "Removal of basic dyes in binary system by adsorption using rarasaponin-bentonite: Revisited of extended Langmuir model", *Chemical Engineering Journal*, 20120501](#)

40

< 1% match (publications)

[Goswami, Aparajita, and Mihir K. Purkait. "The defluoridation of water by acidic alumina", *Chemical Engineering Research and Design*, 2012.](#)

41

< 1% match (student papers from 29-Jun-2015)

[Submitted to Savitribai Phule Pune University on 2015-06-29](#)

paper text:

5Evans blue removal from wastewater by rarasaponin–bentonite I. K. Chandra, Y.-H. Ju, A. Ayucitra & S. Ismadji International Journal of Environmental Science and Technology

ISSN 1735-1472 Volume 10 Number 2 Int. J. Environ. Sci. Technol. (2013) 10:359-370 DOI 10.1007/s13762-012-0114-y

123 Your article is protected by copyright and all rights are held exclusively by CEERS, IAU. This e-offprint is for personal use only and shall not be self-archived in electronic repositories. If you wish to self-archive your work, please use the accepted author's version for posting to your own website or your institution's repository. You may further deposit the accepted author's version on a funder's repository at a funder's request, provided it is not made publicly available until 12 months after publication. 123 DOI 10.1007/s13762 -012- 0114-y ORIGINAL

PAPER

5Evans blue removal from wastewater by rarasaponin–bentonite I. K. Chandra • Y.-H. Ju • A. Ayucitra • S. Ismadji

Received: 2 November 2011 / Revised: 3 January 2012 / Accepted: 25 September 2012 / Published online: 31 October 2012 Ó CEERS, IAU 2012 Abstract The feasibility of natural bentonite and rarasaponin–bentonite for Evans blue removal from aqueous solution was studied. Rarasaponin is a natural surfactant obtained from Sapindus rarak DC was used as modifying agent for natural bentonite modification.

7Adsorption experiments were conducted in a batch system **at various temperatures.** Several **temperature-dependent** isotherm models **(Langmuir, Freundlich, Sips and Toth) were used to** represent **equilibrium data. It was found that** Toth model represents **the**

adsorption

8equilibrium data better than other models.

Kinetic data were best represented

13 **by the pseudo- first order model. The** controlling mechanism of **the adsorption of**

Evans blue onto natural bentonite and rara- saponin–bentonite was physical adsorption. Keywords Bentonite Modification Adsorption Rarasaponin Introduction Many methods have been developed for the purpose of separation or removal of contaminants from industrial wastewater. Among them, adsorption process is still continuously used for sequestering hazardous chemicals from wastewater. In the adsorption process, the correct choice of I. K. Chandra

23 **Y.-H. Ju Department of Chemical Engineering, National Taiwan University of Science and Technology, 43, sec. 4 Keelung Rd., Taipei 106, Taiwan**

I. K. Chandra A. Ayucitra S.

31 **Ismadji (&) Department of Chemical Engineering, Widya Mandala Surabaya Catholic University, Kalijudan 37, Surabaya 60114, Indonesia e-mail: suryadiismadji@yahoo.com**

solid adsorbent is the main key for the success of the process. Activated carbons have been known as superior adsorbents for industrial water or wastewater treatment. However, the main obstacle in using these adsorbents is its price, the commercial activated carbons with high adsorption capacities are expensive, and in the view of economics, the use of these adsorbents for wastewater treatment will significantly increase the production cost of the industry. The search of alternative low-cost adsorbents for wastewater treatment purpose has been begun several decades ago. Mostly the studies focus on biomass-based adsorbents, and some of them focus on clay minerals. Many review papers have summarized the studies on the utilization of biomass-based materials as the

27 **low-cost adsorbents for the removal of** various organics and **heavy metals from** simulated **wastewater**

(Wang and Chen 2006; Volesky 2007; Arief et al. 2008; Lesmana et al. 2009; Febrianto et al. 2009). Indeed, the use of these low-cost adsorbents will reduce the production and operational cost, however for industrial application purpose, the availability of these kinds of adsorbents in large quantity and with constant adsorption capacity is still questionable. Currently, the use of clay minerals for industrial

41 **wastewater treatment applications** is **strongly** recom- mended **due to their local availability, technical feasibility, engineering applications, and cost**

effectiveness.

Among the available clay materials, natural bentonite is the most suitable as the adsorbent for adsorption of pollutants from industrial wastewater. However, because of the hydrophilicity induced by the exchangeable metal cations, natural bentonite usually is not effective in adsorbing organic compounds. The surface of natural bentonite

11 can be changed from hydrophilic to hydrophobic or organophilic by organo-functional molecules such as surface hydroxyl groups, Lewis and Bronsted acidic sites, etc., by grafting organic groups on the clay surface.

For this purpose, the most commonly used chemicals as modifying agents are surfactants (Juang et al. 2002; Lee et al. 2002; Khenifi et al. 2007; Koswojo et al. 2010; Rahardjo et al. 2011). However, most of the surfactants in previous studies are synthetics, which often create serious problem to environment and commonly need expensive waste treatment system for excess or unused surfactants. In relation with environment protection, the study of modification clay using surfactants

33 should be directed towards the use of natural surfactants as modifying agents.

In this study, we employed a natural surfactant called as rarasaponin obtained from the extraction of flesh fruit of Sapindus rarak DC as modifying agent for rarasaponin– bentonite preparation. The natural

39 bentonite used in this study was obtained from a natural bentonite mining located

34 near the border of Pacitan and Ponorogo, East Java, Indonesia. The

adsorption capability of rarasaponin– bentonite was tested by adsorption of dye from synthetic wastewater solution. Evans blue was taken as the dye model. This dye is

15 chemically and photolytically stable, also highly persistent in natural environment due to the presence of

a chromophore group in its molecular structure. It means that the release of this compound in the environment may spread the potential danger of bioaccumulation that may eventually affect human by transport through the food chain (Zee 2002). The adsorption isotherm data obtained in this study were correlated by Langmuir, Freundlich, Sips, and Toth models with its temperature-dependent forms.

Materials and method Materials Dried Sapindus rarak fruit was obtained from Klaten, Central Java, Indonesia. The fruit flesh was separated from the seed and then extracted using deionized water (flesh:solvent = 1:10) at 80 °C for 4 h. The filtrate

30 **was separated from** the **mixture using filter** paper. Subsequently **the** filtrate **was**

evaporated using Buchi RE 121 rotary evaporator and the concentrate was moved into the plastic tubes. The concentrate then dried using Labconco freeze dryer

38 **for 24 h. The dried** rarasaponin **was ground and sieved**

using Retsch Haan screen and the rarasaponin powder with the size of 100/120 mesh was collected and kept in desiccators for further use. The chemical functional groups of rarasaponin powder were determined by Fourier Transform Infrared Spectroscopy (FTIR) and the result is given in Fig. 1. The natural bentonite was obtained from a natural bentonite mining

34 **located near the border of** Pacitan and Ponorogo, **East Java, Indonesia.** Prior to use, **the**

natural bentonite was purified

17 **to remove organic impurities. The** natural bentonite **was**

treated by contacting

30 **with 30 % hydrogen peroxide solution** for 24 h. **The**

excess of

30 **hydrogen peroxide was** then **removed by heating**

the mix- ture at around 100 °C for 1

17 **h. The treated** natural bentonite **was** then **separated from the**

mixture by centrifugation, and

17dried at 110 °C for 24 h.

10Subsequently, the dried natural bentonite was crushed using Janke & Kunkel micro ham- mer mill

and sieved using Retsch Haan screen. The

17chemical composition of the natural bentonite-powder was

analyzed using Rigaku ZSX100e X-Ray Fluorescence. The cation exchange capacity (CEC) of this natural bentonite was analyzed using ASTM C837-99 procedure (63.95 meq/100 g of clay). The diazo dye, Evans blue (C₃₄H₂₄N₆Na₄O₁₄S₄, CAS No. 46160)

11was purchased from Fluka. Analytical grade of hydrogen peroxide was

purchased from Sigma Aldrich. These compounds were used directly without any further purification. Rarasaponin–bentonite preparation Rarasaponin obtained from previous treatment was mixed with natural bentonite powder and dispersed in deionized water and then stirred at 500 rpm for 2 h. The ratio of rarasaponin, natural bentonite, and deionized water was 0.5 g, 5 g, and 25 mL, respectively. Then, the mixture was put into Inextron microwave oven, and heated for 5 min at 700 W. The resultant (rarasaponin–bentonite) was repeatedly washed using distilled water to remove excess rarasaponin. The rarasaponin–bentonite was dried at 100 °C for 24 h. The dried rarasaponin–bentonite was then pulverized until its particle size was around 100/120 mesh. Characterization of adsorbents The natural bentonite and rarasaponin–bentonite was characterized using X-Ray Diffraction (XRD), FTIR, and nitrogen sorption methods. The X-Ray Diffraction analysis was carried out using a Rigaku Miniflex Goniometer instrument using Cu K radiation at 30 kV and 15 mA with

10a step size of 0.01°. The FTIR qualitative analysis

was performed in FTIR Shimadzu 8400s. The method employed for FTIR analysis was KBr technique. To identify the pore structure of both natural bentonite and rarasaponin–bentonite, nitrogen sorption analysis of both samples was carried out in Quadrasorb SI at

35boiling point of nitrogen gas (-196 °C). Before the gas adsorption measurement, the samples were degassed at 150 °C under vacuum condition for 24

h. The nitrogen sorption measurements were conducted at

10 **relative pressure (p/p_0) in the range of 0.**

001–0.998. Adsorption procedure The adsorption isotherm study was performed in batch mode at various temperatures (30, 45, and 60 °C). The known amount of adsorbent was added in a series of Erlenmeyer flasks containing 100 mL of 0.52 mmol/L Evans blue solution. These flasks were then moved into Memmert type WB-14 thermostatic shaker bath and shaken at 100 rpm. The temperature of the thermostatic shaker bath was adjusted to a desired temperature. When equilibrium condition was reached, the adsorbent was separated from the solution by centrifugation (Hettich Zentrifugen EBA-20). The

13 **initial and equilibrium concentration of Evans blue in the solution**

was determined by

39 **Shimadzu UV/VIS-1700 Pharmaspec spectrophotometer at**

its maximum wavelength. The amount of Evans blue adsorbed by the adsorbent was

37 **calculated by the following equation: $q_e = \frac{1}{m} (C_0 - C_e) V$ where q_e is the amount of Evans blue adsorbed**

(mmol/g), m

25 **is the mass of adsorbent (g), while V is the volume of solution (L), and C_0 and C_e are initial and equilibrium concentration (mmol /L), respectively. For the**

kinetic study, the experiments were also conducted isothermally in three different temperatures (30, 45, and 60 °C). A series of Erlenmeyer each containing 100-ml Evans blue solution was mixed with 1 g of adsorbent. The flasks containing the mixtures were then shaken at 100 rpm in Memmert type WB-14 thermostatic shaker bath at certain temperature. At certain interval of time, one of the flasks was taken from the thermostatic shaker bath, and the mixture was separated using centrifugation. The concentration of remaining Evans blue solution was determined using spectrophotometer, and the amount of Evans blue adsorbed at time interval t was determined by the following equation $q_t = \frac{1}{m} (C_0 - C_t) V$ where C_t is the concentration of Evans blue at time interval t . Theory Adsorption is regarded as practical separation method for purification or bulk separation of different kind of industrial products. In the field of environmental pollution control, the adsorption process also plays significant role, especially in the wastewater treatment system. Adsorption equilibrium and kinetic are the most important aspects for the understanding of engineering design methods of adsorption system. The adsorption equilibrium is the most fundamental property and can be correlated through mathematical formulation called as adsorption isotherm. Different kinds of adsorption isotherm models have been developed and proposed. Some of them are based on

simplified physical phenomena of adsorption, while others are purely empirical with two or three empirical parameters. The adsorption of adsorbate onto an adsorbent is affected by temperature, and in most cases the comparisons between the adsorption equilibrium obtained from the experiments and adsorption isotherm model have been made at single temperature. Since the adsorption equilibria are temperature dependent, it is important to include the temperature dependent forms of adsorption models in correlating the experimental data. In this study we used the

7 temperature dependent forms of Langmuir, Freundlich, Sips, and Toth models.

A brief description of the temperature dependent forms of these models are given in this paper, and for complete discussion of these temperature dependent forms the reader can refer to Do (1998). Langmuir isotherm Langmuir (1918) proposed a

32 theory of adsorption on a flat surface

(Do 1998). The

26 theory is based on a kinetic principle, that is the rate of adsorption is equal to the rate of desorption from the surface. The

famous Langmuir model is $q_e = \frac{q_{max} K_L C_e}{1 + K_L C_e}$ where C_e is the equilibrium concentration of the solution (mmol/L) and K_L is a Langmuir affinity constant (L/mmol), q_{max}

38 (mmol/g) is the adsorption capacity

of the adsorbate, while q_e (mmol/g) is the amount of dye adsorbed per unit mass of adsorbent. The temperature dependent forms of the Langmuir isotherm parameters (q_{max} and K_L) can be written in the following forms: $q_{max} = q_{0max} \exp\left[\frac{d}{T_0} \left(\frac{1}{T} - \frac{1}{T_0}\right)\right]$ $K_L = K_{L0} \exp\left[\frac{E}{RT_0} \left(\frac{1}{T} - \frac{1}{T_0}\right)\right]$ where q_{max} is the adsorption capacity at reference temperature T_0 , while the temperature coefficient of the expansion of adsorbate is d . The parameters E and K_{L0} are the heat of adsorption and the Langmuir affinity constant at reference temperature (T_0), respectively. Freundlich isotherm As

12 one of the earliest empirical equations and most widely used isotherm model to describe the adsorption equilibrium data,

Freundlich model (1932) has the form $q_e = K_F C_e^{1/n}$ where K_F (mmol/g)(L/mmol)⁻ⁿ is the measure of Freundlich adsorption capacity and $1/n$ is adsorption intensity. Parameters K_F and n are the complex forms of dependent temperature, and one should not extrapolate them outside their range of validity (Do 1998).

The temperature dependent forms of Freundlich parameters (K_F and n) can be written as $K_F = K_{F0} \exp\left[\frac{a}{T}\right]$

The parameter K_F indicates the Freundlich adsorption capacity at reference temperature (T_0), while a and A_0 are the characteristics of adsorption potential and constant parameter of Clapeyron, respectively. Sips isotherm

14 Sips equation which also known as the Langmuir–Freundlich equation has the form

where K_S (mmol/L) $^{1/n}$ is Sips affinity constant and parameter n is regarded as the parameter characterising the system heterogeneity. When n is unity, the Sips equation reduces to Langmuir, which is suitable for ideal surfaces (Do 1998). Parameters K_S and n have temperature dependent forms as follow: $K_S = K_{S0} \exp \left(\frac{E}{RT} - \frac{E}{T_0} \right)$ $n = n_0 \left(1 - \frac{g}{T} \right)$ The parameters K_{S0} and n_0 are Sips parameters at the reference temperature (T_0), while g is a constant parameter. While Freundlich equation is not valid at low and high concentration and Sips equation also has limitation at low end concentration range, Toth equation gives satisfactory results especially for Henry law type behaviour (Do 1998). This equation is suitable for sub-monolayer coverage system and has the following form: $q_e = \frac{q_{max} C_e^{1-t}}{K_T + C_e^{1-t}}$ Parameter t that is usually less than unity and has the same physical meaning with Sips parameter (n) indicates the system heterogeneity. When its value deviates further away from unity, the system becomes more heterogeneous (Do 1998). Parameter K_T (mmol/L) t is Toth affinity constant. Toth equation also has temperature dependent forms, which is useful to describe adsorption equilibrium data

7 at various temperatures. The temperature dependent forms of

This model can be written as $K_T = K_{T0} \exp \left(\frac{t}{T} - \frac{t}{T_0} \right)$ $t = t_0 \left(1 - \frac{g}{T} \right)$ The affinity coefficient K_{T0} and parameter t_0 are at the reference temperature (T_0), while g is a constant parameter. Theory of adsorption kinetics To properly design an adsorption system, additional information beside the adsorption equilibria is needed. This crucial information is the adsorption kinetic. The rate of pollutants adsorbed into the adsorbent is one of the crucial factors which influence the effectiveness of the sorption process (Plazinski et al. 2009). Many mathematical models have been proposed, but

13 pseudo- first order and pseudo-second order are still extensively applied to describe the kinetics

of sorption in solid/ solution systems. Table 1 The FTIR spectra of rarasaponin Functional group Wavenumber (1/cm)

36 O–H stretch, free hydroxyl **C–H stretch** **C=O stretch**

(ester carbonyls group) C=C stretch (olefin group) C–CH₃ bend C–H bend bonded with hydroxyl group C–O stretch (carbonyls group) C=C stretch (ethers group) 3,580.60 2,928.71 1,729.06 1,645.17 1,447.48 1,380.94 1,248.82 1,048.24 Pseudo-first order Results and discussion

9This model proposed by Lagergren in nineteenth century (Lagergren 1898), and the pseudo-first order equation

Identification of rarasaponin became popular

9to describe the rate of sorption in the liquid-phase systems. The differential form of

the pseudo- The FTIR spectrum of rarasaponin obtained from extrac- first model is as follows: tion of Sapindus rarak is given in Fig. 1; Table 1. The hydroxyl group of rarasaponin was shown at wavenumber $3,580.60 \text{ cm}^{-1}$, while the wavenumber $1,729.06 \text{ cm}^{-1}$ indicates ester carbonyl group. Olefin group and ethers The integration form of Eq. (15) can be written as group were identified at wavenumber $1,645.17$ and follows: $1,048.24 \text{ cm}^{-1}$, respectively. These functional groups indicates that the structure of rarasaponin belongs to oli- $\frac{1}{4} \exp(-k_1 t)$ gosaccharide (Asao et al. 2009). Pseudo-second order Characterization of adsorbents Another model which is widely used for interpretation of The chemical composition of natural bentonite and rara- adsorption kinetic data is the pseudo-second order. The saponin–bentonite were determined by X-ray fluorescence

9pseudo-second order kinetic is usually associated with the method and the

results are given in Table 2. The chemical

9situation when the rate of direct adsorption/desorption

composition of rarasaponin–bentonite in general is slightly process controls the overall sorption kinetics (Plazinski different from the natural one. The change of the chemical et al. 2009). This model was first developed by Blanchard composition of the adsorbents is possibly due to the et al. (1984) and the linearized form of this model proposed by Ho (1995). The mathematical form of the pseudo-second order model is $\frac{dq}{dt} = k_2(q - q_0)^2$ Table 2 The XRF analysis of natural bentonite and rarasaponin– bentonite Compound Percentage (%) By integrating of Eq. (17) the final form of the pseudo-second order is Natural bentonite Rarasaponin–bentonite Al_2O_3 15.90 16.30 $\frac{1}{4} \exp(-k_2 t)$ SiO_2 49.80 49.10 Fe_2O_3 7.94 7.72 Both pseudo-first and pseudo-second order equations CaO 2.36 1.83 have parameter time constant k (g/mmol min) to describe K_2O 0.93 0.87 the rate constant of adsorption, the symbol q_t (mmol/g) Na_2O 0.08 0.09 represents the amount of adsorbate on the surface of the MgO 2.41 2.89 adsorbent at any time, t (Plazinski et al. 2009). Other 20.58 21.20

10Fig. 2 FTIR spectra of natural bentonite and rarasaponin –bentonite attachment of the

functional groups from rarasaponin structure on the surface of bentonite. The FTIR-spectra result (Fig. 2) showed that there are some functional groups either appear or lost after the modification process. The spectral bands were observed at 673.11 and 674.07

21cm-1, respectively, for natural bentonite and rarasaponin –bentonite,

which represent the Al

21–O–Si groups of the octahedral sheet, while the **Si–O– Si**

bond on tetrahedral sheet appears at 443.60

21cm-1 for natural bentonite and

shifting to 421.42 cm⁻¹ for rarasaponin–bentonite. For rarasaponin–bentonite, a spectral band observed at 1,251.72 cm⁻¹ indicates the C=O stretch of deacylated group from rarasaponin. This evidence shows that the interaction between rarasaponin and natural bentonite occurred during modification process. For easy reference, the FTIR spectra of both adsorbents are also tabulated in Table 3. Figure 3 depicts the XRD results of natural bentonite and rarasaponin–bentonite. The basal spacing d₀₀₁ for natural bentonite and rarasaponin–bentonite is 14.9940 and 16.6202 Å at 5.8895 (2θ) and 5.3128 (2θ), respectively. From these data, montmorillonite was identified as major component from natural bentonite (Tabak et al. 2007). This basal spacing was expanded after modification process with rarasaponin possibly due to intercalation of rarasaponin molecules into bentonite interlayer spaces (Kurniawan et al. 2011). The expansion of natural bentonite interlamellar spacing will affect surface characteristic of rarasaponin–bentonite, including polarizability, electronic charge, and pore dimension (Do 1998). Table 3

10FTIR spectra of natural bentonite and rarasaponin –bentonite

Functional group Wavenumber (1/cm) Natural Rarasaponin– bentonite bentonite Al(Mg)–O–H stretching H–O–H stretching (for H₂O) O–H stretching of silanol (Si–OH) groups H–O–H bending C=O stretch of deacylated carbonyl group Si–O–Si stretching Si–O stretching O–H bending bounded 2Al³⁺ O–H bending bounded Mg²⁺ and Al³⁺ O–H bending bounded Fe³⁺ and Al³⁺ Si–O stretching of silica and quartz Al–O–Si bending (for octahedral Al) Si–O–Si bending 3,637.50 3,496.70 3,268.15 1,650.95 – – 1,106.10 929.63 842.83 879.48 803.30 673.11 443.60 3,626.89 3,438.84 3,271.05 1,650.95 1,251.72 1,014.49 1,105.14 919.98 838.01 – 792.69 674.07 421.42 Fig. 3 XRD results of natural bentonite and rarasaponin–bentonite Adsorption study Several adsorption isotherm models which initially developed

14for gas-phase adsorption can be used to

represent the liquid-phase adsorption experimental data (Rahardjo et al. 2011).

20In this study, four **isotherm models** (Langmuir, **Freundlich**, **Sips**, and **Toth**)

14with their **temperature** dependent **forms** were **employed to correlate the adsorption experimental data of**

Evans blue onto natural and rarasa-ponin-bentonite. A non-linear least square method was employed to obtain the parameters of the adsorption models. To obtain the best fitted parameters of each model, the

5sum of squared error **was** employed **as objective function to be minimized.**

"P SSE $\frac{1}{n} \sum_{i=1}^n (q_{e,exp} - q_{e,cal})^2$ where $q_{e,exp}$ (mmol/g) and $q_{e,cal}$ (mmol/g) are the actual amount and the calculated value of dye adsorbed by the adsorbent, respectively. While n is the total number of experimental data used. Figures 4 and 5 show the adsorption equilibrium isotherms of Evans blue on natural bentonite and rarasa-ponin-bentonite at various temperatures, and the fits of 0,04 different adsorption isotherm models. In these figures, the isotherm fittings are represented as solid lines while

4the **experimental data are** given **as symbols.**

It seems that all the models can represent the adsorption

4experimental data well. The optimal parameters from the fitting of Langmuir,

Freundlich, Sips, and Toth equations

4with the adsorption **experimental data are** summarized **in Table**

4. Even visually all the models can represent the

4data well as seen in Figs.

4 and 5, however, the decision to choose the correct isotherm should not be based just only on how good the model represents the data visually or the value of SSE. The decision should be based on the physical meaning of each fitted parameter obtained. If the values of fitted parameters of the model are reasonable and consistent with the physical meaning of the parameter, it means that the model is applicable and can be used to represent the experimental data. 0,04 0,03 q_e (mmol/g) 0,02 0,01 $T = 303K$ $T = 318K$ $T = 333K$ 0,00 0,0 0,1 0,2 0,3 0,4 0,5

6Ce (mmol/L) (a) 0,04 0,03 qe (mmol/g) 0,02 0,

01 T = 303 K T = 318 K T = 333 K

30,00 0,0 0,1 0,2 0,3 0,

4 0,5

6Ce (mmol/L) (b) 0,04 0,03 qe (mmol/g) 0,02 0,

01 T = 303K T = 318K T = 333K 0,00 0,0 0,1 0,2 0,3 0,4 0,5 Ce (mmol/L) (c)

16qe (mmol/g) 0,03 0,02 0,01 0,00 0,0 0,1 0,2 0,3 Ce (mmol/L)

(d) T = 303K T = 318K T = 333K 0,4 0,5 Fig. 4 Adsorption experimental data of Evans blue onto natural bentonite and the model fitted by

20a Langmuir, b Freundlich, c Sips, and d Toth

isotherms 0,08 0,08 0,06 0,06

6qe (mmol/g) 0,04 qe (mmol/g) 0,04 0,02 0,00

T = 303K T = 318K T = 333K 0,0 0,1 0,2 0,3 0,4

36Ce (mmol/L) (a) 0,02 0,00 0,0 0,1 0,

2 Ce (mmol/L) (b) 0,3

3T = 303 K T = 318 K T = 333 K 0,4 0,08 0,08 0,

06

16qe (mmol/g) 0,04 0,02 0,00 0,0 0,1 0,2 Ce (mmol/L)

(c)

60,06 qe (mmol/g) 0,04 T = 303K 0,02

T = 318K T = 333K 0,00 0,3 0,4 0,0 0,1 0,2 Ce (mmol/L) (d) T = 303K T = 318K T = 333K 0,3 0,4 Fig. 5
Adsorption experimental data of Evans blue onto rarasaponin–bentonite and the model fitted by

20a Langmuir, b Freundlich, c Sips, and d Toth

isotherms Since all the models used in this study can represent the experimental data well with small values of SSE, we will discuss further about the consistency of the physical meaning of each parameters listed in Table 4. The adsorption capacity of the adsorbent is given by parameter q_{max} for Langmuir, Sips and Toth, and K_F for Freundlich. The parameter of the adsorption capacity for all models, both for natural bentonite and rarasaponin-bentonite, is reasonable and the value is consistent with literatures as indicated in Table 4. The Langmuir, Sips, and Toth models have the parameter d which represents the temperature coefficient of expansion of adsorbate. As mentioned by Ismadji and Bhatia (2001) this parameter is specific for a given component and independent of the type of adsorbent. From Table 4 it can be seen that only Toth equation gave consistent fitted parameter d . While the fitted parameter d for Langmuir and Sips is not consistent and dependent on the type of adsorbent. So, essentially Langmuir and Sips models fail to correlate the adsorption data of Evans blue onto natural bentonite and rarasaponin–bentonite. As mentioned previously, K_L (Langmuir), k_0 (Sips), and k_T (Toth) are affinity parameters. These parameter mea-

12how strong an adsorbate molecule is attracted onto a surface. When the affinity parameter is high, the surface is covered with more adsorbate molecules as a result of stronger affinity towards the surface of

adsorbent (Do 1998). Figures 4 and 5 clearly indicate that rarasaponin– bentonite has higher adsorption capacity than its parent form. This is a strong indication that Evans blue has stronger adsorption affinity towards the surface of rarasaponin–bentonite than natural bentonite. Inconsistency of the fitted affinity parameter value was observed for Langmuir and Sips as seen in Table 4. Since Langmuir and Sips Table 4 The fitted temperature dependent parameters of several isotherm models Isotherm Parameters Natural bentonite Organo-bentonite Langmuir Freundlich Sips Toth q_{max} (mmol/g) d 9 102 (K⁻¹) K_L (L/mmol) E (kJ/mol) SSE k_0 (mmol/g)(mmol/L)⁻ⁿ A_0 SSE q_{max} (mmol/g) d 9 102 (K⁻¹) K_S (L/mmol) E (kJ/mol) n_0 g SSE q_{max} (mmol/g) d 9 102 (K⁻¹) K_T (L/mmol) E (kJ/mol) t_0 g SSE 0.167 0.340 0.452 0.88 0.06 0.274 1.646 3.121 0.03 0.185 0.810 0.542 5.29 1.021 0.299 0.03 0.263 2.380 2.051 9.63 0.779 3.383 0.07 0.527 0.070 0.114 3.15 3.34 0.612 1.457 2.858 0.05 0.575 1.330 0.367 9.64 1.008 0.405 0.05 0.516 2.230 2.524 11.43 0.978 4.243 0.09 models fail to give reasonable and consistent parameters (temperature coefficient of expansion of adsorbate and adsorption affinity), both of these models will not be included in subsequent discussion. The heterogeneity of the system is given by parameter t_0 for Toth equation and parameter A_0 for Freundlich model. As mentioned by Do (1998) that the system heterogeneity could stem

32from the solid or the adsorbate or a combination of both.

During the modification process, the acyl (C₂H₃O?) groups which are attached to the carbonyl group of the rarasaponin molecules were deacylated and attached to the protonated silanol groups which are available in tetrahedral sheet of natural bentonite (Kurniawan et al. 2011), the

33attachment of the rarasaponin molecules into interlayer structure of natural bentonite

increases the het- erogeneity of the system. This phenomenon was not cap- tured by parameter A₀ in Freundlich model, the value of this parameter decreases with increase of the system het- erogeneity. The value of parameter t₀ in Toth equation increases with increase of the system heterogeneity as indicated in Table 4. Therefore, only Toth model still can represent the adsorption experimental data.

27In order to determine the applicability of Toth equation to represent the adsorption experimental data

of Evans blue onto natural bentonite and rarasaponin–bentonite we still need to examine the rest of fitted parameters (E and g). In the physical

15adsorption, the temperature has been known to have negative effect on the amount of adsorbate adsorbed by the solid. The uptake of adsorbate decreases with increase

of temperature. Physical adsorption processes usually have isosteric heat of adsorption ≈ 40 kJ/mol (Do 1998). Figures 4 and 5 show that the adsorption of Evans blue onto natural bentonite and rarasaponin–bentonite is mainly controlled by physical adsorption. In the physical adsorption, the increase of temperature weakens the inter- action between Evans blue and natural bentonite or rara- saponin–bentonite, therefore the amount of dye uptake by both of the adsorbents decreased with increasing temper- ature. The results of parameter E

4from the fitting of Toth equation is consistent with our experimental data as depicted in

Figs. 4 and 5. Higher rate of heat adsorption on rarasaponin–bentonite is an indication that other bonding mechanism also took place during the process. As men- tioned in the previous paragraph that the acyl groups in the rarasaponin structure play significant role during the for- mation of rarasaponin-bentonite. Some of these acyl groups were deacylated during the formation of rarasaponin–ben- tonite and some of them remain in the carbonyl group of 0,30 0

29,25 0,20 Q_e (mmol/g) 0,15 0,10

3T = 303 K 0,05 T = 318 K T = 333 K 0, 00 0

20 40 60 80 100 120 140 t (minutes) (i-a) 0

18,5 0,4 Q_e (mmol/g) 0,3 0,2 0,1

3T = 303 K T = 318 K T = 333 K 0,0 0

20 40 60 80 100 120 140 t (minutes) (ii-a) 0,30 0

29,25 0,20 Q_e (mmol/g) 0,15 0,10

3T = 303 K 0,05 T = 318 K T = 333 K 0, 00 0

20 40 60 80 100 120 140 t (minutes) (i-b) 0

18,5 0,4 Q_e (mmol/g) 0,3 0,2 0,1

3T = 303 K T = 318 K T = 333 K 0,0 0

20 40 60 80 100 120 140 t (minutes) (ii-b) Fig. 6 Kinetics experimental data of Evans blue onto natural bentonite and fitted model by (i-a) pseudo-first order, (i-b) pseudo-second order models, and using rarasaponin–bentonite (ii-a)

28pseudo-first order, (ii-b) pseudo-second order models Table 5 kinetic parameters for

4pseudo-first and pseudo-second order models

T (K)

22 Pseudo-first order Pseudo-second order k1 (min-1) qe (mmol/g) R2 SSE k2 (g.mmol-1 min-1) qe (mmol/g) R2

SSE Using natural bentonite as adsorbent 303 0.1491 0.2683 0.9578 318 0.1107 0.2499 0.9609 333 0.0945 0.2399 0.9574 Using rarasaponin–bentonite as adsorbent 303 0.1403 0.4393 0.9607 318 0.1223 0.4063 0.9704 333 0.0964 0.3599 0.9575 0.0215 0.6593 0.0201 0.4721 0.0208 0.3946 0.0336 0.3873 0.0272 0.3379 0.0303 0.2678 0.2929 0.9204 0.2791 0.9316 0.2719 0.9267 0.4785 0.9188 0.4494 0.9474 0.4082 0.9388 0.0295 0.0266 0.0272 0.0483 0.0362 0.0364 rarasaponin structure. In polar solutions such as water, excess acyl groups in the surface of rarasaponin–bentonite became positively charged. On the other side, the Evans blue which belongs to Azo dye category was negatively charged due to its SO₃ functional groups. The electrostatic interaction between positive charge of acyl groups and negative charge of Evans blue occurred, leading to higher rate of heat adsorption. In the Toth equation, parameter g also measures the heterogeneity of the system. If the value of this parameter deviates from unity, the system is more heterogeneous. The value of fitted parameter g of Toth model increases with system heterogeneity as shown in Table 4.

8 Based on the evaluation of the physical meaning of fitted parameters of each equation, it is clear that the Toth equation can represent the adsorption data better than the other equations. Adsorption kinetics One of

the most crucial factors for designing the adsorption system is the ability to predict the rate at which the adsorbate removal takes place in a given solid/solution system (Plazinski et al. 2009). Numerous kinetic models have been developed to predict the behavior of the adsorption kinetic experimental data, and most of the models were developed based on certain fundamental approach to interfacial kinetics (Plazinski et al. 2009). Most widely used

40 models such as pseudo-first order (Lagergren 1898) and pseudo-second order (Blanchard et al. 1984)

were also developed based on the interfacial kinetics approach. Figure 6 presents the adsorption kinetic experimental data Evans blue and the calculated values using

24 pseudo- first order and pseudo-second order models for Natural bentonite and rarasaponin–bentonite, respectively. The fit- ted parameters'

value of

24 pseudo-first order and pseudo- second order are summarized in **Table 5**. Both of **the** models can represent **the**

4 **experimental data well. The pseudo-first order**

gave a smaller SSE a slightly better coefficient of correlation (R^2). The deviation of q_e obtained from the fitting and

28 **experimental data in the pseudo first order** is smaller than **the pseudo-second order.**

Based on this evidence, the controlling mechanism of the adsorption of Evans blue onto natural bentonite and rarasaponin– bentonite was physical adsorption Conclusion

8 **The adsorption of Evans blue onto natural bentonite and rarasaponin –bentonite was studied.**

The modification of natural bentonite using natural surfactant (rarasaponin) increased the adsorption capacity of the bentonite. Four adsorption isotherm models with their temperature dependent forms were used to correlate the adsorption experimental data, and it was found that the Toth model gave the consistent and reasonable values of fitted parameters. For the kinetic study, the pseudo-first order gave smaller deviation of the q_e value obtained from the fitting with the experimental data. The controlling mechanism of the adsorption of Evans blue onto natural bentonite and rarasaponin–bentonite was physical adsorption, other mechanism such as electrostatic interaction also occurred. The best way to determine the adsorption mechanism is using direct spectroscopic measurement such as calorimeter, etc.

Acknowledgments The first author

19 **would like to express their sincere gratitude to the Department of Chemical Engineering National Taiwan University of Science and**

Technology, Taiwan, for undergraduate

19 **exchange student grant and providing all the research facilities.**

References Arief VO, Trilestari K, Sunarso J, Indraswati N, Ismadji S (2008) Recent progress on biosorption of heavy metals from liquids using low cost biosorbents: characterization, biosorption parameters and mechanism studies. *Clean Soil Air Water* 36:937–962 Asao Y, Morikawa T, Xie Y, Okamoto M, Hamao M, Matsuda H, Muraoka O, Yuan D, Yoshikawa M (2009) Structures of acetylated oleanane-type triterpene saponins, rarasaponins IV, V, and VI, and anti-hyperlipidemic constituents from the pericarps of *Sapindus rarak*. *Chem Pharm Bull* 57:198–203 Blanchard G, Maunaye M, Martin G (1984) Removal of heavy metals from waters by means of natural zeolites. *Water Res* 18(12):1501–1507 Do DD (1998) Adsorption analysis :

equilibria and kinetics. Imperial College Press, Australia Febrianto J, Kosasih AN, Sunarso J, Ju YH, Indraswati N, Ismadji S (2009) Equilibrium and kinetic studies in adsorption of heavy metals using biosorbent: a summary of recent studies. *J Hazard Mater* 162:616–645 Ho YS (1995) Adsorption of heavy metals from waste streams by peat. PhD Thesis, The University of Birmingham, Birmingham Ismadji S, Bhatia SK (2001) A modified pore filling isotherm for liquid phase adsorption in activated carbon. *Langmuir* 17:1488–1498 Juang RS, Lin S-H, Tsao K-H (2002) Mechanism of sorption of phenols from aqueous solutions onto surfactant-modified montmorillonite. *J Colloid Interface Sci* 254:234–241 Khenifi A, Bouberka Z, Sekrane F, Kameche M, Derriche Z (2007) Adsorption study of an industrial dye by an organic clay. *Adsorption* 13:149–158 Koswojo E, Utomo RP, Ju YH, Ayucitra A, Soetaredjo FE, Sunarso J, Ismadji S (2010) Acid green 25 removal from wastewater by organo-bentonite from Pacitan. *Appl Clay Sci* 48:81–86 Kurniawan A, Sutiono H, Ju YH, Soetaredjo FE, Ayucitra A, Yudha A, Ismadji S (2011) Utilization of rarasaponin natural surfactant for organo-bentonite preparation: application for methylene blue removal from aqueous effluent. *Microporous Mesoporous Mater* 142:184–193 Lagergren S (1898) Zur theorie der sogenannten adsorption gelöster stoffe *Kungliga Svenska Vetenskapsakademiens Handlingar* 24:1–39 Lee JJ, Choi J, Park JW (2002) Simultaneous sorption of lead and chlorobenzene by organobentonite. *Chemosphere* 49:1309–1315 Lesmana SO, Febriana N, Soetaredjo FE, Sunarso J, Ismadji S (2009) Studies on potential applications of biomass for the separation of heavy metals from water and wastewater. *Biochem Eng J* 44:19–41 Plazinski W, Rudzinski W, Plazinska A (2009) Theoretical models of sorption kinetics including a surface reaction mechanism: a review. *Adv Colloid Interface Sci* 152:2–13 Rahardjo AK, Susanto MJJ, Kurniawan A, Indraswati N, Ismadji S Volesky B (2007) Biosorption and me. *Water Res* 41:4017–4029 (2011) Modified Ponorogo bentonite for the removal of ampicillin. *Wang J, Chen C (2006) Biosorption of heavy metals by Saccharomyces cerevisiae: a review. Biotechnol Adv* 24:427–451 Tabak A, Afsin B, Aygun S, F Koksal E (2007) Structural Zee FP (2002) Anaerobic azo dyes reduction. Doctoral Thesis, characteristics of organo-modified bentonites of different origin. Wageningen University, Wageningen, The Netherlands *J Thermal Anal Calor* 87:375–381

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

360 Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

362 Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

363 364 Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

365 366 Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

367 368 Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

Author's personal copy

2Int. J. Environ. Sci. Technol. (2013) 10 :359–370

369 370 Author's personal copy Int. J. Environ. Sci. Technol. (2013) 10:359–370 123 123 123 123 123 123
123 123 123 123 123 123